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Paper: CHEMISTRY	OBJECTIVE TYPE	Class: 10th
Time: 15 Min.	(2nd 50%)	Max. Marks: 12
Name: _____		Dated: _____

Note: Give answers to the questions on the objective answer sheet provided. Four possible answers A, B, C or D are given for each question. Fill the circle (A, B, C or D) which you consider correct with marker or pen.

Q.1 **[1 × 12= 12]**

- Which is not a heavy metal?
(a) Cadmium (b) Lead (c) Zinc (d) Mercury
- Which of the following is a triglyceride?
(a) Carbohydrates (b) Proteins (c) Lipids (d) Vitamins
- Percentage of carbon dioxide in dry air is
(a) 21% (b) 0.93% (c) 0.03% (d) None of these
- Thousands of the amino acid polymerize to form
(a) carbohydrates (b) proteins (c) lipids (d) vitamins
- The percentage of Sunlight absorbed by atmosphere gases is
(a) 2% (b) 10% (c) 18% (d) 25%
- Which gas acts as a glass wall of a green house?
(a) Oxygen (b) Carbon Dioxides (c) Sulphur Dioxide (d) Hydrogen
- Which one of the following salts makes the water permanently hard?
(a) Na_2CO_3 (b) NaHCO_3 (c) $\text{Ca}(\text{HCO}_3)_2$ (d) CaSO_4
- Which one of the following vitamins is water soluble?
(a) vitamin A (b) vitamin E (c) vitamin C (d) vitamin D
- Water has a maximum density at
(a) 10°C (b) 0°C (c) 4°C (d) 100°C
- Which contains sufficient amount of metal
(a) Mineral (b) Ores (c) Rocks (d) Soil
- The process by which atmospheric nitrogen is turned into nitrates in the soil is called:
(a) nitration (b) fixing (c) oxidation (d) redaction
- Which one of the followings is not a fraction of petroleum?
(a) kerosene oil (b) diesel oil (c) alcohol (d) petrol

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Test: Chemistry	GROUP	Class: 10th
Time: 1 Hour.	Date:	Max. Marks: 30
Name: _____		Section:

Q.No.1. Tick the right answer.**(1 x 6)**

- i) Concentration of copper ore is carried out by:
 a) Calcinations b) Roasting c) Forth flotation d) Distillation
- ii) Which one of the following is not a fraction of petroleum:
 a) Kerosene oil b) Diesel c) Alcohol d) Petrol
- iii) At the time of independence, there were big industrial units in India:
 a) 921 b) 34 c) 821 d) 43
- iv) Formula of copper glance is:
 a) CuFeS_2 b) Cu_2S c) CuFeS d) CuS
- v) ICI was established in
 a) 1844 b) 1944 c) 1866 d) 1966
- vi) Which is used for dry cleansing purposes?
 a) Petroleum b) Petroleum gas c) Diesel oil d) Fuel oil

Q.No.2 Give Short Answers.**(2 x 6)**

- What is metallurgy? Write names of step of metallurgy.
- What is the principle of Solvay's Process?
- What is petroleum?
- What is blister copper?
- What is difference between diesel oil and fuel oil?
- Describe the principle of gravity separation to concentrate metal ore

Q.No.3 a) How urea is increased in manufactured. Explain showing the flow sheet?**(04)****b) Explain the process of smelting with reference to copper.****(03)****ATP****Q.No.4 Write the procedure and observation to demonstrate that sugar decomposes into elements or other compounds. (05)****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.com@gmail.com

Test: Chemistry	GROUP	Class: 10th
Time: 1 Hour.	Date:	Max. Marks: 30
Name: _____		Section:

Q.No.1. Tick the right answer.**(1 x 6)**

- i) It is used as jet fuel:
 a) Fuel oil b) Diesel oil c) Kerosene oil d) CNG
- ii) Matte is a mixture of:
 a) FeS and CuS b) Cu_2O and FeO c) Cu_2S and FeS d) CuS and FeO
- iii) Formula of urea is:
 a) $\text{NH}_2\text{COONH}_2$ b) NH_2CONH_2 c) $\text{NH}_2\text{COONH}_4$ d) NH_2CONH_4
- iv) At the time of independence how many big industrial units came to the share of Pakistan:
 a) 921 b) 219 c) 34 d) 43
- v) Which is used as flux in smelting process?
 a) Coke b) Sand c) Lime stone d) None of these
- vi) Sindh alkalies limited was establish:
 a) 1966 b) 1944 c) 1866 d) 1844

Q.No.2 Give Short Answers.**(2 x 6)**

- What is an ore? Give example
- What are the advantages of Solvay's Process?
- How petroleum is extracted from underground deposits?
- How ammonia is recovered in Solvay's Process?
- What is difference between crude oil and residual oil?
- What the principle is of forth flotation process to concentrate metal ore?

Q.No. 3 a) Write a note on fractional distillation of petroleum**(04)****b) Explain the process of Bessemerization with reference to copper.****(03)****ATP****Q.No. 4 Write the observations of the experiments done during the softening of hard water****(05)**

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Test: Chemistry	GROUP	Class: 10th
Time: 40 min.	Date:	Max. Marks: 30
Name: _____		Section:

Q.No.1. Tick the right answer.**(1 x 6)**

- i) Which one of the following compound is found in kerosene oil:
 a) C₁₀H₂₂ b) C₁₁H₂₄ c) C₁₂H₂₆ d) All of these
- ii) Urea is used in automobile system to reduce pollutants in exhaust gases:
 a) N₂ b) SO₂ c) NO_x d) CO₂
- iii) ICI is located in:
 a) Karachi b) Khewra c) Kalabagh d) Abbotabad
- iv) Crude oil is heated in the furnace up to:
 a) 200°C b) 300°C c) 400°C d) 500°C
- v) Oil used in forth flotation process is:
 a) Kerosene oil b) Mustard oil c) Coconut oil d) Pine oil
- vi) Colour of hair is caused by the presence of compounds of:
 a) Alkali metal b) Alkaline earth metals c) Transition metals d) Non metals

Q. No.2 Give Short Answers.**(2 x 7)**

- What is difference between mineral and ore?
- Why flux is added during the extraction of metals?
- Describe Haber process to prepare ammonia?
- How carbon dioxide is prepared in Solvay's Process?
- What do you mean by fractional distillation?
- How can we put out oil fires?
- What do you mean by fraction of petroleum? Write names of fractions of petroleum?

Q. No. 3a) What is blister copper? How it will be purified?**(04)****b) What are natural fertilizers? Write their advantages.****(03)**ATP**Q. No. 4 Why hardness of water containing CaCl₂ is not removed by adding NaHCO₃ until it is heated?****(03)****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.com@gmail.com

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Test: Chemistry	Group-	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section: _____

Q1. Tick the correct answer from the following. (4x1=04)

- Molecular formula of hexane is.
(a) C_6H_{14} (b) C_7H_{16} (c) C_8H_{18} (d) None of these
- _____ have fruity smell.
(a) Ethers (b) Alcohols (c) Esters (d) Ketones
- Formula of form aldehyde is
(a) $CH_3 - \overset{\overset{O}{||}}{C} - H$ (b) $CH - \overset{\overset{O}{||}}{C} - H$ (c) $CH_3 - \overset{\overset{O}{||}}{C} - O$ (d) CH_3COH
- Each person's sweat contains a unique blend of _____.
(a) Alcohol (b) Ketones (c) Carboxylic acid (d) Amines

Q2. Answer the following short questions. 2x4=08

- What are alkenes? Write their general formula and example.
- Describe different radicals of butane.
- What are ketones? Write their general formula.
- Write the tests for esters.

Q.3. Long Question.

- What are the characteristics of homologous series? **04**
- Describe the functional group of alcohols. How alcoholic groups are identified? **04**

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Test: Chemistry	Group-	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section: _____

Q1. Tick the correct answer from the following. (4x1=04)

- General formula of alkyl radicals is.
(a) C_nH_{2n+2} (b) C_nH_{2n-2} (c) C_nH_{2n+1} (d) C_nH_{2n}
- In which group oxygen is attached on both sides with carbon atoms.
(a) Ketones (b) Ethers (c) Aldehydes (d) Esters
- General formula of alkenes is.
(a) C_nH_{2n} (b) C_nH_{2n+2} (c) C_nH_{2n+1} (d) C_nH_{2n-2}
- Identify which one of the following compounds is a ketone
(a) $(CH_3)_2CH_3OH$ (b) $(CH_3)_2CO$ (c) $(CH_3)_2NH$ (d) $(CH_3)_2CHCL$

Q2. Answer the following short questions. 2x4=08

- What are alkyl radicals? Write their general formula.
- Draw the alkyl radicals of propane.
- Describe the functional group of aldehydes.
- How can you test ketones.

Q.3. Long Question.

- What is a functional group? Explain difference between aldehydic and ketonic groups. How both are identified from each other. **04**
- What are amines? Explain different types of amines giving an example of each type. How primary amino group is identified. **04**

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Paper: CHEMISTRY		Class: 10th
Time: 1 Hour 45 Minutes.	SUBJECTIVE TYPE	Max. Marks: 48
Name: _____	(2 nd 50%)	Dated: _____

PART I

Note: Out of Q. No. 2, 3 & 4. Attempt any five questions from each section. Short answer must contain two sentences at least. [15 × 2 = 30]

(Section - A)**Q.2. Give short answers of the followings: (any five)**

- Differentiate between primary and secondary pollutants.
- What is the function of DNA?
- How margarine is produced?
- How proteins are formed?
- Describe the sources and uses of vitamin A.
- Name of major constituents of troposphere.
- How plants synthesize carbohydrates?
- How acid rain increase the acidity of soil?

(Section - B)**Q.3. Give short answers of the followings: (any five)**

- How aquatic life is affected by acid rain?
- Differentiate between soft and hard water.
- CO is a hidden enemy explains its action.
- Why the flood risks are increasing?
- What is the function of fertilizers?
- Why CO₂ gas is called a green house gas?
- How water rises in plants?
- How lime water is dissolves in water?

(Section - C)**Q.4. Give short answers of the followings: (any five)**

- Why the water molecule is Polar?
- What do you mean by boiler scales? How are they removed?
- How pesticides cause water pollution?
- Give the advantages of Solvay's process.
- What is the difference between diesel oil and fuel oil?
- How ammonia is recovered in the Solvay's process?
- How NaHCO₃ is converted to Na₂CO₃?
- Define metallurgy.

PART II**Note: Attempt any two questions.****[2 x 9 = 18]****Q.5. (a) Explain the source and uses of proteins.****[5]****(b) Write the effects of ozone depletion.****[4]****Q.6. (a) Define acid rain. How it forms and what are its effects?****[5]****(b) Explain important waterborne diseases. How can these be prevented?****[4]****Q.7. (a) Explain the methods for removing permanent hardness of water.****[5]****(b) Write a note in detail on smelting and bessemerization, giving a specific example.****[4]**

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Paper: CHEMISTRY	OBJECTIVE TYPE	Class: 10th
Time: 15 Min.	(2nd 50%)	Max. Marks: 12
Name: _____		Dated: _____

Note: Give answers to the questions on the objective answer sheet provided. Four possible answers A, B, C or D are given for each question. Fill the circle (A, B, C or D) which you consider correct with marker or pen.

Q.1 **[1 × 12= 12]**

1. Which is not a heavy metal?
(a) Cadmium (b) Lead (c) Zinc (d) Mercury
2. Which of the following is a triglyceride?
(a) Carbohydrates (b) Proteins (c) Lipids (d) Vitamins
3. Percentage of carbon dioxide in dry air is
(a) 21% (b) 0.93% (c) 0.03% (d) None of these
4. Thousands of the amino acid polymerize to form
(a) carbohydrates (b) proteins (c) lipids (d) vitamins
5. The percentage of Sunlight absorbed by atmosphere gases is
(a) 2% (b) 10% (c) 18% (d) 25%
6. Which gas acts as a glass wall of a green house?
(a) Oxygen (b) Carbon Dioxides (c) Sulphur Dioxide (d) Hydrogen
7. Which one of the following salts makes the water permanently hard?
(a) Na₂CO₃ (b) NaHCO₃ (c) Ca(HCO₃)₂ (d) CaSO₄
8. Which one of the following vitamins is water soluble?
(a) vitamin A (b) vitamin E (c) vitamin C (d) vitamin D
9. Water has a maximum density at
(a) 10°C (b) 0°C (c) 4°C (d) 100°C
10. Which contains sufficient amount of metal
(a) Mineral (b) Ores (c) Rocks (d) Soil
11. The process by which atmospheric nitrogen is turned into nitrates in the soil is called:
(a) nitration (b) fixing (c) oxidation (d) redaction
12. Which one of the followings is not a fraction of petroleum?
(a) kerosene oil (b) diesel oil (c) alcohol (d) petrol

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Paper: CHEMISTRY	SUBJECTIVE TYPE	Class: 10th
Time: 1 Hour 45 Minutes.	(2 nd 50%)	Max. Marks: 38
Name: _____		Dated: _____

PART I

Note: Out of Q. No. 2, 3 & 4. Attempt any five questions from each section. Short answer must contain two sentences at least. [19 × 2 = 38]

(Section - A)

Q.2. Give short answers of the followings: (any six)

- Differentiate between primary and secondary pollutants.
- What is the function of DNA?
- How margarine is produced?
- How proteins are formed?
- Describe the sources and uses of vitamin A.
- Name of major constituents of troposphere.
- How plants synthesize carbohydrates?
- How acid rain increase the acidity of soil?

(Section - B)

Q.3. Give short answers of the followings: (any six)

- How aquatic life is affected by acid rain?
- Differentiate between soft and hard water.
- CO is a hidden enemy explains its action.
- Why the flood risks are increasing?
- What is the function of fertilizers?
- Why CO₂ gas is called a green house gas?
- How water rises in plants?
- How lime water is dissolves in water?

(Section - C)

Q.4. Give short answers of the followings: (any seven)

- Why the water molecule is Polar?
- What do you mean by boiler scales? How are they removed?
- How pesticides cause water pollution?
- Give the advantages of Solvay's process.
- What is the difference between diesel oil and fuel oil?
- How ammonia is recovered in the Solvay's process?
- How NaHCO₃ is converted to Na₂CO₃?
- Define metallurgy.

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Paper: CHEMISTRY	OBJECTIVE TYPE	Class: 10th
Time: 15 Min.	(2nd 50%)	Max. Marks: 12
Name: _____		Dated: _____

Note: Give answers to the questions on the objective answer sheet provided. Four possible answers A, B, C or D are given for each question. Fill the circle (A, B, C or D) which you consider correct with marker or pen.

Q.1 **[1 × 12= 12]**

1. Which is not a heavy metal?
(a) Cadmium (b) Lead (c) Zinc (d) Mercury
2. Which of the following is a triglyceride?
(a) Carbohydrates (b) Proteins (c) Lipids (d) Vitamins
3. Percentage of carbon dioxide in dry air is
(a) 21% (b) 0.93% (c) 0.03% (d) None of these
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6. Which gas acts as a glass wall of a green house?
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7. Which one of the following salts makes the water permanently hard?
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8. Which one of the following vitamins is water soluble?
(a) vitamin A (b) vitamin E (c) vitamin C (d) vitamin D
9. Water has a maximum density at
(a) 10°C (b) 0°C (c) 4°C (d) 100°C
10. Which contains sufficient amount of metal
(a) Mineral (b) Ores (c) Rocks (d) Soil
11. The process by which atmospheric nitrogen is turned into nitrates in the soil is called:
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12. Which one of the followings is not a fraction of petroleum?
(a) kerosene oil (b) diesel oil (c) alcohol (d) petrol

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Paper: CHEMISTRY	First Term Examination (50%)	Class: 10th
Time: 15 Min.	OBJECTIVE TYPE	Max. Marks: 12
Name: _____		

Paper Code	0	0	0	0
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Note: Give answers to the questions on the objective answer sheet provided. Four possible answers A, B, C or D are given for each question. Fill the circle (A, B, C or D) which you consider correct with marker or pen.

Q.1 **[12 x 1 = 12]**

Q. No.	Question	(A)	(B)	(C)	(D)
1.	The two major components of atmosphere are	Carbon and nitrogen	Nitrogen and oxygen	Oxygen and chlorine	None of these
2.	Which one of the following is a Lewis base?	NH ₃	BF ₃	H ⁺	AlCl ₃
3.	You want to dry a gas which one of the following salt you will use?	CaCl ₂	NaCl	CaO	Na ₂ SiO ₃
4.	Which of the following is used as foaming agent in fire extinguishers?	Magnesium hydroxide	Aluminium hydroxide	Potassium hydroxide	Calcium hydroxide
5.	The percentage of Carbon in peat is	60%	70%	80%	90%
6.	Alkyl radical is derivative of:	Alkane	Alkene	Alkyne	All of these
7.	Ester is recognized by their.	yellow colour	Deep colour	Unpleasant odour	Fruity smell
8.	Carbon tetrachloride is used in	dry cleaning	waxes	rubber	None of these
9.	Dehydration of alcohols can be carried out with	NaOH	KOH	H ₂ SO ₄	HCl
10.	The colour of hydrogen iodide is	Purple	Black	Red	Colourless
11.	Malic acid is present in.	Apple	Fats	Sting of bees	Urine
12.	Which one of the following does not contain protein?	Pulses	Potatoes	Beans	Eggs

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Paper: CHEMISTRY

First Term Examination (50%)

Class: 10th

Time: 2 Hours 45 Minutes.

SUBJECTIVE TYPE

Max. Marks: 63

Name: _____

PART I**Q.2. Write short answers of the following (ANY FIVE) (5×2=10)**

- What are irreversible reactions? Give few characteristics of them.
- Why equilibrium state is attainable from either way?
- How dynamic equilibrium is established?
- Write the names of two chemicals in which nitrogen is used.
- What do you mean by the extent of reaction?
- Derive equilibrium constant expression for the synthesis of ammonia from nitrogen and hydrogen.
- Define Arrhenius acid. Give example.
- Define adduct. Give example.

Q.3. Write short answers of the following (ANY SIX) (6×2=12)

- Define complex salt. Give example.
- Why BF_3 behave as a Lewis acid?
- Write the characteristics of Lewis acid.
- Prove that water is amphoteric specie.
- Who rejected the vital force theory and how?
- What are two basic conditions for elements to exhibit catenation?
- Define alkyl radical. How they are formed?
- Define functional group. Give example.
- Define isomerism. Give example.

Q.4. Write short answers of the following (ANY FIVE) (5×2=10)

- Why C – C bonds are much stronger than the Si – Si bonds?
- What is petroleum? Give its composition.
- How alkyl halides are reduced?
- Give the physical properties of alkenes.
- Why colour of bromine water discharge on addition of ethene?
- How tetrabromo ethane prepared from acetylene?
- Why the burning of alkanes required sufficient supply of oxygen?
- Draw the structural formula of iso butane and iso pentane.

PART II**Note: ATTEMPT ANY THREE QUESTIONS (7×3=21)**

- Q.5. (a) State the law of mass action and derive the expression for equilibrium constant for a general reaction. [4]
(b) Explain Lewis concept of acid and base with example. [3]
- Q.6. (a) What is auto-ionization of water? How it is used to establish the pH of water? [4]
(b) Define homologous series. Write down the characteristics of homologous series. [3]
- Q.7. (a) Write the uses of ethene. [4]
(b) Write the preparation of alkanes. [3]
- Q.8. (a) Write down the macroscopic characteristics of dynamic equilibrium. [4]
(b) What are normal and basic salts? Give examples. [3]
- Q.9. (a) Write the general characteristics of organic compounds. [4]
(b) Explain the oxidation of acetylene. [3]

A.T.P.**Note: Attempt any TWO questions. 2×5=10**

- Q.10 (a) Which material is required to standardize the given NaOH volumetrically. (2)
(b) How can you demonstrate that some natural substances are weak acids? (3)
- Q.11 (a) Write the observations (2)

Metallic ion	Sodium ion	Calcium ion	Copper ion	Barium ion
Colour imparted to the flame				

- (b) Write the procedure to identify sodium, calcium, strontium, barium and copper ions by flame test. (3)
- Q.12 (a) Which material is required to determine the exact molarity of a solution of oxalic acid? (2)
(b) Write the procedure to determine the exact molarity of Na_2CO_3 solution volumetrically. (3)

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Q.1

[12 x 1 = 12]

Q. No.	Question	(A)	(B)	(C)	(D)
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SEDIINFO.NET SCIENCE ACADEMYsedinfo.net@gmail.com**609-B Faisal Town Lahore.****Paper: CHEMISTRY****Class: 10th****Time: 2 Hours 45 Minutes.****SUBJECTIVE TYPE****Max. Marks: 63****Name: _____****PART I****Q.2. Write short answers of the following (ANY FIVE)****(5×2=10)**

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- Q.5. (a) State the law of mass action and derive the expression for equilibrium constant for a general reaction. [4]
(b) Explain Lewis concept of acid and base with example. [3]
- Q.6. (a) What is auto-ionization of water? How it is used to establish the pH of water? [4]
(b) Define homologous series. Write down the characteristics of homologous series. [3]
- Q.7. (a) Write the uses of ethene. [4]
(b) Write the preparation of alkanes. [3]
- Q.8. (a) Write down the macroscopic characteristics of dynamic equilibrium. [4]
(b) What are normal and basic salts? Give examples. [3]
- Q.9. (a) Write the general characteristics of organic compounds. [4]
(b) Explain the oxidation of acetylene. [3]

A.T.P.**Note: Attempt any TWO questions.****2×5=10**

- Q.10 (a) Which material is required to standardize the given NaOH volumetrically? (2)
(b) How can you demonstrate that some natural substances are weak acids? (3)
- Q.11 (a) Write the observations (2)

Metallic ion	Sodium ion	Calcium ion	Copper ion	Barium ion
Colour imported to the flame				

- (b) Write the procedure to identify sodium, calcium, strontium, barium and copper ions by flame test. (3)
- Q.12 (a) Which material is required to determine the exact molarity of a solution of oxalic acid? (2)
(b) Write the procedure to determine the exact molarity of Na_2CO_3 solution volumetrically. (3)

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Test: Chemistry Time: 40 Min. Name: _____	Group Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Normally ran water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas
- Infrared radiation emitted by the earth are absorbed by
a) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Earth has natural systems.
a) one b) two c) three d) four

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Why CO₂ gas is called green house gas?
- How aquatic life is affected by acid rain
- Write the composition of dry air.
- Point out the two serious effects of ozone depletion.
- What is incineration?
- State the major source of CO and CO₂.
- Explain the phenomenon of decreasing temperature in troposphere.
- Define atmosphere.

Q. No.3.

- How acid rain forms? What are its effects? 5
- How ozone layer is depleted? 4

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Test: Chemistry Time: 40 Min. Name: _____	Group Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Earth has natural systems.
a) one b) two c) three d) four
- Infrared radiation emitted by the earth are absorbed by
b) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Normally ran water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 6 = 16**

- Name the major constituents of troposphere.
- What do you mean by air pollutant?
- Why flood risk increases?
- CO₂ is hidden enemy, Explain.
- What are the threats are to human health to SO₂ gas as air pollutant?
- Hoe acid rain increases the acidity of soil?
- What do you mean by ozone hole?
- How ozone layer is formed?

Q. No.3.

- How acid rain forms? What are its effects? 5
- How ozone layer is depleted? 4

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Test: Chemistry Time: 40 Min. Name: _____	Class: 10th Max. Marks: 30 Section: _____
Date: _____	

Q. No.1. Choose the Correct Answer.

1 x 7 = 7

1. When a reaction ceases to proceed further, it is called
 a) Chemical state b) static state c) physical state d) none of these
2. The % age of nitrogen and oxygen in our atmosphere is
 a) 98% b) 90% c) 85% d) 99%
3. The oxidation of carbon monoxide goes to completion at 1000 K the value of K_c is
 a) 3.0×10^{-9} b) 2.2×10^{22} c) 1.8×10^{-19} d) 1.1×10^{24}
4. For the reaction of PCl₃ and Cl₂ to form PCl₅ the unit of K_c are
 a) mol dm⁻³ b) mol⁻¹ dm⁻³ c) mol⁻¹ dm³ d) mol dm³
5. The reaction which have comparable amount of reactant and products at equilibrium state have
 a) Very small K_c value b) moderate K_c value
 c) very long K_c value d) none of these
6. SO₂ gas used to manufacturing of
 a) H₂SO₄ b) HNO₃ c) HCl d) H₂CO₃
7. The value of K_c depends only on
 a) pressure b) temperature c) atmosphere d) none of these

Q. No.2. Give Short Answers of the following questions.

2 x 7 = 14

1. Give the characteristics of reversible reaction.
2. Why at equilibrium state reaction does not stop?
3. How direct of reaction can be predicted?
4. Write the equilibrium constant expression for following equations
 a) $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$ b) $\text{CO} + 3\text{H}_2 \rightleftharpoons \text{CH}_4 + \text{H}_2\text{O}$
5. If the reaction has larger value of K_c will it go to completion and why?
6. What do you mean by equilibrium constant?
7. What is static equilibrium? Explain with example.

Q. No.3. a) State the law of mass action and derive the expression for equilibrium constant for a general reaction.

5

b) Write the observations.

4

Metallic ion	Sodium ion	Calcium ion	Barium ion	Copper ion
Colour imparted to the flame				

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.

1 x 7 = 7

1. SO₂ gas used to manufacturing of
a) H₂SO₄ b) HNO₃ c) HCl d) H₂CO₃
2. The value of K_c depends only on
a) pressure b) temperature c) atmosphere d) none of these
3. The reaction which have comparable amount of reactant and products at equilibrium state have
a) Very small K_c value b) moderate K_c value
c) very long K_c value d) none of these
4. For the reaction of PCl₃ and Cl₂ to form PCl₅ the unit of K_c are
a) mol dm⁻³ b) mol⁻¹ dm⁻³ c) mol⁻¹ dm³ d) mol dm³
5. The oxidation of carbon monoxide goes to completion at 1000 K the value of K_c is
a) 3.0 x 10⁻⁹ b) 2.2 x 10²² c) 1.8 x 10⁻¹⁹ d) 1.1 x 10²⁴
6. When a reaction ceases to proceed further, it is called
a) Chemical state b) static state c) physical state d) none of these
7. The % age of nitrogen and oxygen in our atmosphere is
a) 98% b) 90% c) 85% d) 99%

Q. No.2. Give Short Answers of the following questions.

2 x 7 = 14

1. Give the characteristics of reversible reaction.
2. Why at equilibrium state reaction does not stop?
3. How direct of reaction can be predicted?
4. Write the equilibrium constant expression for following equations
b) $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$ b) $\text{CO} + 3\text{H}_2\text{O} \rightleftharpoons \text{CH}_4 + \text{H}_2\text{O}$
5. If the reaction has larger value of K_c will it go to completion and why?
6. What do you mean by equilibrium constant?
7. What is static equilibrium? Explain with example.

Q. No.3. a) Write the macroscopic characteristics of dynamic equilibrium.

5

b) Write the procedure to standardize the given NaOH solution volumetrically.

4

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Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- The formula of phosphoric acid is
a) H_2PO_4 b) H_3PO_4 c) H_2SO_3 d) CH_3COOH
- Which of the following will use Natural food preservatives are:
a) Salts b) Sugar c) Alcohols d) all of these
- Which one of the following salt is used in manufacture of flint glass?
a) CaO b) $CaCl_2$ c) $NaClO_3$ d) KNO_3
- Which of the following is neutral salt?
a) $NaHPO_4$ b) $Zn(OH)NO_3$ c) KCl d) none of these
- Formula of potash alum is
a) $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24 H_2O$ b) $KClO_3$
c) $CuSO_4 \cdot 5H_2O$ d) none of these
- Example of mixed salt
a) $Pb(CH_3COO)_2$ b) $Ca(OCl)Cl$ c) NH_4NO_3 d) $KHSO_4$
- Which one is insoluble salt?
a) $NaCl$ b) $CuSO_4$ c) $CuCO_3$ d) Na_2CO_3

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define pH. What is pH of water?
- How insoluble salts are prepared?
- Why H^+ ion acts as a Lewis acid?
- How the basic salts turn into normal salts?
- Give the limitations of Arrhenius concept.
- Give the uses of calcium oxide.
- What are complex salts?

Q. No.3. a) Write the procedure to standardize the give HCl solution volumetrically.**5****b) Define acid and base according to Lowry-Bronsted concept. Explain with examples. 4****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Which acid is used for etching designs on copper plates?
a) H_2SO_4 b) HNO_3 c) HCl d) CH_3COOH
- You want to dry a gas which of the following will use Natural food preservatives are:
a) Salts b) Sugar c) Alcohols d) all of these
- When acid react with carbonates and bicarbonates which gas is evolved?
a) H_2 b) CO_2 c) Cl_2 d) N_2
- Which one of the following salt is used in manufacture of flint glass?
a) CaO b) $CaCl_2$ c) $NaClO_3$ d) KNO_3
- Example of mixed salt
a) $Pb(CH_3COO)_2$ b) $Ca(OCl)Cl$ c) NH_4NO_3 d) $KHSO_4$
- Which one is insoluble salt?
a) $NaCl$ b) $CuSO_4$ c) $CuCO_3$ d) Na_2CO_3
- The formula of phosphoric acid is
a) H_2PO_4 b) H_3PO_4 c) H_2SO_3 d) CH_3COOH

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- How insoluble salts are prepared?
- Give few characteristics of salts.
- Explain neutralization reaction according to Lewis concept.
- What are complex salts?
- How double salts are prepared?
- Prove that water is amphoteric compound.
- Why BF_3 behave as a Lewis acid?

Q. No.3. a) What is auto ionization? How it is used to establish the pH of water?**5****b) Write the observations to determine the exact molarity of Na_2CO_3 solution volumetrically. 4**

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Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Which is proteins
a) Pulses b) beans c) butter d) both a and b
- Which into is absorbed directly small intestine into blood stream?
a) Starch b) lipide c) fats d) glucose
- Smell of methyl butonoate is like
a) apple b) pineapple c) orange d) both a and b
- Deficiency of vitamin D causes.
a) Night blindness b) Rickets c) Cancer d) none of these
- Which of the following is reducing sugar?
a) glucose b) maltose c) fructose d) starch
- When glucose and fructose combine they produce
a) Starch b) cellulose c) vitamin d) none of these
- The organic compound used to control bleeding are
a) Vitamin b) proteins c) lipids d) glycerides

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define lipids. Give its general formula.
- How margarine is formed?
- Which elements are found in proteins?
- Why RNA are called messenger?
- How protein is formed?
- What is the significance of vitamins?
- How gelatin is obtained?

Q. No.3. a) Write the commercial used of microbial and lactase enzymes.**5****b) How can you identify aldehydes using Fehling solution?****4****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- The most important oligosaccharide is
a) glucose b) Fructose c) maltose d) sucrose
- Which one is triglyceride?
a) Portions b) lipids c) vitamins d) DNA
- Which of the following vitamins is water soluble?
a) vitamins A b) vitamins B c) vitamins C d) vitamins D
- Smell of ethyl butanoate is like
a) apple b) pine apple c) orange d) foul
- Which of the following is high energy food
a) lipids b) proteins c) carbohydrates d) none
- Carbohydrates provide energy per gram.
a) 17 KJ b) 170 KJ c) 200 KJ d) 1700 KJ
- Formula of palmitic acid is
a) $C_{12}H_{22}O_{11}$ b) $C_{15}H_{31}COOH$ c) $(NH_2)_2CO$ d) $C_{17}H_{35}COOH$

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Give the characteristics of monosaccharide's.
- What is the function of DNA?
- What is difference between glucose and fructose?
- What is the importance and source of vitamin A?
- How proteins are formed?
- What are the components of nucleotide?
- Write the used of amylase enzymes.

Q. No.3. a) Write the sources and uses of proteins?**5****b) How can you identify ketones using 2-4 dinitrophenyl hydrazine test?****4**

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Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Which is proteins
a) Pulses b) beans c) butter d) both a and b
- Which into is absorbed directly small intestine into blood stream?
a) Starch b) lipide c) fats d) glucose
- Smell of methyl butonoate is like
a) apple b) pineapple c) orange d) both a and b
- Deficiency of vitamin D causes.
a) Night blindness b) Rickets c) Cancer d) none of these
- Which of the following is reducing sugar?
a) glucose b) maltose c) fructose d) starch
- When glucose and fructose combine they produce
a) Starch b) cellulose c) vitamin d) none of these
- The organic compound used to control bleeding are
a) Vitamin b) proteins c) lipids d) glycerides

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define lipids. Give its general formula.
- How margarine is formed?
- Which elements are found in proteins?
- Why RNA are called messenger?
- How protein is formed?
- What is the significance of vitamins?
- How gelatin is obtained?

Q. No.3. a) Write the commercial used of microbial and lactase enzymes.**5****b) How can you identify aldehydes using Fehling solution?****4****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- The most important oligosaccharide is
a) glucose b) Fructose c) maltose d) sucrose
- Which one is triglyceride?
a) Portions b) lipids c) vitamins d) DNA
- Which of the following vitamins is water soluble?
a) vitamins A b) vitamins B c) vitamins C d) vitamins D
- Smell of ethyl butanoate is like
a) apple b) pine apple c) orange d) foul
- Which of the following Is high energy food
a) lipids b) proteins c) carbohydrates d) none
- Carbohydrates provide energy per gram.
a) 17 KJ b) 170 KJ c) 200 KJ d) 1700 KJ
- Formula of palmitic acid is
a) $C_{12}H_{22}O_{11}$ b) $C_{15}H_{31}COOH$ c) $(NH_2)_2CO$ d) $C_{17}H_{35}COOH$

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Give the characteristics of monosaccharide's.
- What is the function of DNA?
- What is difference between glucose and fructose?
- What is the importance and source of vitamin A?
- How proteins are formed?
- What are the components of nucleotide?
- Write the used of amylase enzymes.

Q. No.3. a) Write the sources and uses of proteins?**5****b) How can you identify ketones using 2-4 dinitrophenyl hydrazine test?****4**

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Test: Chemistry	Class: 10 th
Time: 40 Min.	Max. Marks: 20
Name: _____	Section: _____
Group- Date:	

Q. No.1. Choose the Correct Answer.**1 x 4 = 04**

- The reactions, which proceed in both ways are called
a) Reversible b) irreversible c) Addition d) none
- Color of iodine is
a) Red b) purple c) yellow d) green
- For the reaction of PCl_3 and Cl_2 to form PCl_5 , the unit of K_C are
a) mol.dm^{-3} b) $\text{mol}^{-1}.\text{dm}^{-3}$ c) $\text{mol}^{-1}.\text{dm}^3$ d) mol.dm^3
- At initial stage the rate of forward reaction is
a) Very slow b) very fast c) both a and b d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What are irreversible reactions? Give its few characteristics.
- Define reversible reaction.
- Why at equilibrium state reaction does not stop?
- Define chemical equilibrium state.
- Define static equilibrium? Give example.

Q. No.3. State the law of mass action. Drive the expression for equilibrium constant for a general equation.**06****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10 th
Time: 40 Min.	Max. Marks: 20
Name: _____	Section: _____
Group- Date:	

Q. No.1. Choose the Correct Answer.**1 x 4 = 04**

- The reactions, which proceed in both ways are called
a) Reversible b) irreversible c) Addition d) none
- Color of iodine is
a) Red b) purple c) yellow d) green
- For the reaction of PCl_3 and Cl_2 to form PCl_5 , the unit of K_C are
a) mol.dm^{-3} b) $\text{mol}^{-1}.\text{dm}^{-3}$ c) $\text{mol}^{-1}.\text{dm}^3$ d) mol.dm^3
- At initial stage the rate of forward reaction is
a) Very slow b) very fast c) both a and b d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define reversible reaction.
- Define static equilibrium? Give example.
- Why at equilibrium state reaction does not stop?
- What are irreversible reactions? Give its few characteristics.
- Define chemical equilibrium state.

Q. No.3. Write the macroscopic characteristics of dynamic equilibriums.**06**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Night blindness is because of deficiency of
 - Vitamin A
 - Vitamin B
 - Vitamin C
 - Vitamin D
- Which one of the following is a triglyceride?
 - Carbohydrates
 - proteins
 - lipids
 - Vitamins
- Animal fats are used in
 - Medicine
 - bakery items
 - Soap industry
 - none
- The comical formula of stearic acid is
 - $C_{15}H_{31}COOH$
 - $C_{15}H_{35}COOH$
 - $C_{17}H_{35}COOH$
 - none
- Rancid butter has a foul smell because of
 - Malic acid
 - butanoic acid
 - palmitic acid
 - none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Plants are source of oils, Justify?
- What is the function of DNA?
- Write the characteristics of poly saccharine.
- Describe the source and uses of vitamin A.
- Why RNA is called a messenger?

Q. No.3. Write sources and uses of carbohydrates.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Night blindness is because of deficiency of
 - Vitamin A
 - Vitamin B
 - Vitamin C
 - Vitamin D
- Which one of the following is a triglyceride?
 - Carbohydrates
 - proteins
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 - Vitamins
- Animal fats are used in
 - Medicine
 - bakery items
 - Soap industry
 - none
- The comical formula of stearic acid is
 - $C_{15}H_{31}COOH$
 - $C_{15}H_{35}COOH$
 - $C_{17}H_{35}COOH$
 - none
- Rancid butter has a foul smell because of
 - Malic acid
 - butanoic acid
 - palmitic acid
 - none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define carbohydrates. Give its general formula.
- What is the function of DNA?
- Give the general formula of lipids.
- Describe the source and uses of vitamin A.
- Why RNA is called a messenger?

Q. No.3. Write note on oligosaccharides.**5**

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Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of the following is a triglyceride?
a) Carbohydrates b) proteins c) lipids d) Vitamins
- Animal fats are used in
a) Medicine b) bakery items c) Soap industry d) none
- The chemical formula of stearic acid is
a) $C_{15}H_{31}COOH$ b) $C_{15}H_{35}COOH$ c) $C_{17}H_{35}COOH$ d) none
- Rancid butter has a foul smell because of
a) Malic acid b) butanoic acid c) palmitic acid d) none
- Which one is tasteless?
(a) Starch (b) Glucose (c) Fructose (d) Sucrose

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Plants are source of oils, Justify?
- How proteins are formed?
- Write the character of polysaccharides.
- How gelatin is obtained?
- How plants synthesized carbohydrates?

Q. No.3. Write sources and uses of carbohydrates.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of the following is a triglyceride?
a) Carbohydrates b) proteins c) lipids d) Vitamins
- Animal fats are used in
a) Medicine b) bakery items c) Soap industry d) none
- The chemical formula of stearic acid is
a) $C_{15}H_{31}COOH$ b) $C_{15}H_{35}COOH$ c) $C_{17}H_{35}COOH$ d) none
- Rancid butter has a foul smell because of
a) Malic acid b) butanoic acid c) palmitic acid d) none
- The most important oligosaccharide is.
(a) Maltose (b) Glucose (c) Fructose (d) Sucrose

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define carbohydrates. Give its general formula.
- What do you mean by non-essential amino acid?
- Give the general formula of lipids.
- Draw the structure of glucose.
- What is dextrose? Give its function.

Q. No.3. Write note on oligosaccharides.**5**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of the following is a triglyceride?
a) Carbohydrates b) proteins c) lipids d) Vitamins
- Animal fats are used in
a) Medicine b) bakery items c) Soap industry d) none
- The chemical formula of stearic acid is
a) $C_{15}H_{31}COOH$ b) $C_{15}H_{35}COOH$ c) $C_{17}H_{35}COOH$ d) none
- Rancid butter has a foul smell because of
a) Malic acid b) butanoic acid c) palmitic acid d) none
- Which one is tasteless?
(a) Starch (b) Glucose (c) Fructose (d) Sucrose

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Plants are source of oils, Justify?
- How proteins are formed?
- Write the character of polysaccharides.
- How gelatin is obtained?
- How plants synthesized carbohydrates?

Q. No.3. Write sources and uses of carbohydrates.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of the following is a triglyceride?
a) Carbohydrates b) proteins c) lipids d) Vitamins
- Animal fats are used in
a) Medicine b) bakery items c) Soap industry d) none
- The chemical formula of stearic acid is
a) $C_{15}H_{31}COOH$ b) $C_{15}H_{35}COOH$ c) $C_{17}H_{35}COOH$ d) none
- Rancid butter has a foul smell because of
a) Malic acid b) butanoic acid c) palmitic acid d) none
- The most important oligosaccharide is.
(a) Maltose (b) Glucose (c) Fructose (d) Sucrose

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define carbohydrates. Give its general formula.
- What do you mean by non-essential amino acid?
- Give the general formula of lipids.
- Draw the structure of glucose.
- What is dextrose? Give its function.

Q. No.3. Write note on oligosaccharides.**5**

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Test: Chemistry	Group-	Class: 10th
Time: 45 mins	Date:	Max. Marks: 20
Name: _____		Section:

Q 1. Tick the correct options. (1x5=5)

- Acidus means:
 - Sour
 - Bitter
 - Saltish
 - None
- Ca(OH)₂ is a base according to:
 - Arrhenius
 - Bronsted Lowry
 - Lewis
 - None
- BF₃ is an acid according to:
 - Lewis
 - Bronsted Lowry
 - Arrhenius
 - None
- H₃O⁺ is:
 - Conjugate acid
 - Conjugate base
 - Amphoteric
 - None
- Arrhenius Concept was given in:
 - 1786
 - 1787
 - 1923
 - 1920

Q.No.2 Answer the following questions: (2x5=10)

- What is the concept of Sir Humphery Davy about acids?
- Give any three characteristics of Acids.
- Write the limitations of Bronsted Lowry Concept of acids and base.
- Define Adduct and give its example.
- Prove that H₂O is Amphoteric compound.

Q 3. Long Questions. (05)

Discuss Arrhenius Acids and Bases in detail by giving various examples?

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Test: Chemistry	Group-	Class: 10th
Time: 45 mins	Date:	Max. Marks: 20
Name: _____		Section:

Q.No.1 . Tick the correct options. (1x5=5)

- CN⁻ acts as:
 - Lewis base
 - Lewis acid
 - Arrhenius acids
 - All
- According to Lewis H⁺ behaves:
 - Base
 - Acid
 - Neutral
 - None
- Base turns litmus:
 - Red to Blue
 - Blue to Red
 - Both a & b
 - None
- CH₃COO⁻ is:
 - Conjugate acid
 - Conjugate base
 - Both a & b
 - None
- Na⁺ behaves as Lewis:
 - Acid
 - Base
 - Both a & b
 - None

Q.No.2 Answer the following questions: (2x5=10)

- Define adduct?
- Why BF₃ behaves as Lewis acid?
- Define conjugate base.
- Give the limitations of Bronsted Lowry concept of acids and bases.
- Write any three characteristics of bases.

Q.No.3 Long Questions. (05)

Discuss Bronsted Lowry Concept of Acids and bases in detail.

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Test: Chemistry	Group-	Class: 10th
Time: 45 mins	Date:	Max. Marks: 20
Name: _____		Section:

Q.No.1 . Tick the correct options. (1x5=5)

1. According to Arrhenius CH_3COOH is a:
a) Acid b) Base c) Both a & b d) None
2. ROH behaves as:
a) Lewis acid b) Lewis base c) Arrhenius acid d) None
3. Jabir Bin Hayan prepared:
a) CH_3COOH b) H_3PO_4 c) H_2PO_4 d) None
4. Lewis concept was presented in:
a) 1923 b) 1924 c) 1922 d) None
5. The bond formed between Lewis acid and base is:
a) Ionic b) Covalent c) Metallic d) Co-ordinate covalent bond

Q.No.2 Answer the following questions: (2x5=10)

- i) What is the concept of Sir Humphery Davy about acids?
- ii) Give any three characteristics of bases?
- vi) Give the limitations of Arrhenius Theory?
- vii) Why water is Amphoteric species?
- viii) Define conjugate base. Give example.

Q.No.3 Long Questions. (05)

Discuss Lewis Concept of Acids and bases with examples.

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Test: Chemistry	Prof. Aamir Abbas	Class: 10th
Time: 45 mins	Group-	Max. Marks: 20
Name: _____	Date:	Section:

Q 1. Tick the correct options. (1x5=5)

- Lactic Acid is found in:
 - Apples
 - Fats
 - Urine
 - Sour Milk
- The colour of Zn(OH)₂ ppt is:
 - Blue
 - Brown
 - White
 - None
- The value of [H⁺] is at 25°C:
 - 1.0×10⁻⁶μ
 - 1.0×10⁻⁷ mol dm⁻³
 - 1.5×10⁻⁷ mol dm⁻³
 - None
- Acidic solution having pH value:
 - Less than 7
 - More than 7
 - Equal to 7
 - None
- The chemical formula of Potash Alum having water molecules:
 - 24
 - 22
 - 20
 - 14

Q.No.2 Answer the following questions: (2x5=10)

- Give the uses of magnesium hydroxide?
- Define indicators and give the example?
- How the salts are named?
- Why HCl and H₂SO₄ have different pH values?
- How acids react with metallic oxide?

Q 3. Long Questions.

What is auto-ionization of water? How it is used to establishes pH of water? (05)

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Test: Chemistry	Prof. Aamir Abbas	Class: 10th
Time: 45 mins	Group-	Max. Marks: 20
Name: _____	Date:	Section:

Q.No.1 . Tick the correct options. (1x5=5)

- Which acid is found in Rancid butter:
 - Citric acid
 - Lactic acid
 - Butyric acid
 - Uric acid
- The colour of Fe(OH)₂ ppt is:
 - Blue
 - Brown
 - White
 - Dirty green
- Water is an electrolyte:
 - Strong
 - Weak
 - Non electrolyte
 - None
- Basic solution having pH value:
 - More than 7
 - Less than 7
 - Equal to 7
 - All of these
- Mohr's salt having water molecules:
 - 6
 - 24
 - 5
 - None

Q.No.2 Answer the following questions: (2x5=10)

- Give the uses of calcium oxide.
- What are the uses of pH?
- Define universal indicator?
- How insoluble salts are prepared?
- How will you justify that salts are neutral compounds?

Q.No.3 Long Questions.

(05)

Write a note on Normal Salts; Basic Salts; complex salts

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Test: Chemistry	Group-C	Class: 10th
Time: 45 mins	Date:	Max. Marks: 20
Name: _____		Section:

Q 1. Tick the correct options. (1x5=5)

- Acidus means:
 - Sour
 - Bitter
 - Saltish
 - None
- Ca(OH)₂ is a base according to:
 - Arrhenius
 - Bronsted Lowry
 - Lewis
 - None
- BF₃ is an acid according to:
 - Lewis
 - Bronsted Lowry
 - Arrhenius
 - None
- H₃O⁺ is:
 - Conjugate acid
 - Conjugate base
 - Amphoteric
 - None
- Arrhenius Concept was given in:
 - 1786
 - 1787
 - 1923
 - 1920

Q.No.2 Answer the following questions: (2x5=10)

- What is the concept of Sir Humphery Davy about acids?
- Give any three characteristics of Acids.
- Write the limitations of Bronsted Lowry Concept of acids and base.
- Define Adduct and give its example.
- Prove that H₂O is Amphoteric compound.

Q 3. Long Questions. (05)

Discuss Arrhenius Acids and Bases in detail by giving various examples?

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Test: Chemistry	Group-	Class: 10th
Time: 45 mins	Date:	Max. Marks: 20
Name: _____		Section:

Q.No.1 . Tick the correct options. (1x5=5)

- CN⁻ acts as:
 - Lewis base
 - Lewis acid
 - Arrhenius acids
 - All
- According to Lewis H⁺ behaves:
 - Base
 - Acid
 - Neutral
 - None
- Base turns litmus:
 - Red to Blue
 - Blue to Red
 - Both a & b
 - None
- CH₃COO⁻ is:
 - Conjugate acid
 - Conjugate base
 - Both a & b
 - None
- Na⁺ behaves as Lewis:
 - Acid
 - Base
 - Both a & b
 - None

Q.No.2 Answer the following questions: (2x5=10)

- Define adduct?
- Why BF₃ behaves as Lewis acid?
- Define conjugate base.
- Give the limitations of Bronsted Lowry concept of acids and bases.
- Write any three characteristics of bases.

Q.No.3 Long Questions. (05)

Discuss Bronsted Lowry Concept of Acids and bases in detail.

SEDIINFO.NET SCIENCE ACADEMYsedinfo.net@gmail.com**Test: Chemistry****Time: 45 mins****Name: _____****Group-****Date:****Class: 10th****Max. Marks: 20****Section:****Q.No.1 . Tick the correct options.****(1x5=5)**

1. According to Arrhenius CH_3COOH is a:
a) Acid b) Base c) Both a & b d) None
2. ROH behaves as:
a) Lewis acid b) Lewis base c) Arrhenius acid d) None
3. Jabir Bin Hayan prepared:
a) CH_3COOH b) H_3PO_4 c) H_2PO_4 d) None
4. Lewis concept was presented in:
a) 1923 b) 1924 c) 1922 d) None
5. The bond formed between Lewis acid and base is:
a) Ionic b) Covalent c) Metallic d) Co-ordinate covalent bond

Q.No.2 Answer the following questions:**(2x5=10)**

- i) What is the concept of Sir Humphery Davy about acids?
- ii) Give any three characteristics of bases?
- vi) Give the limitations of Arrhenius Theory?
- vii) Why water is Amphoteric species?
- viii) Define conjugate base. Give example.

Q.No.3 Long Questions.**(05)**

Discuss Lewis Concept of Acids and bases with examples.

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Test: Chemistry	Group-	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q1. Answer the following Short Question.**5x1=05**

- Sodium hydroxide.
 - Plaster of Paris
 - drying
 - soap
 - paper
- pH value of 0.04M HCl solution is
 - 2
 - 11
 - 3.5
 - 1.7
- Fats contains
 - uric acid
 - malic acid
 - citric acid
 - Formic acid
- Which is act as Lewis base?
 - Cl⁻
 - H⁺
 - N₂⁺
 - none
- What is the pOH of a 0.02 M Ca(OH)₂
 - 1.397
 - 1.698
 - 12.31
 - 12.61

Q. No. 2. Give Answers of the following short questions:**5x2=10**

- Write the uses of sulphuric acid.
- Write the physical properties of acid.
- How the insoluble salts are prepared?
- Define pH. What is the pH of bases?
- Define indicators. Give example.

Q. No. 3. Long Question:

- What is auto-ionization of water? How is it used to establish the pH of water? **05**

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Test: Chemistry	Group-	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q1. Answer the following Short Question.**5x1=05**

- Potassium hydroxide is used in:
 - Soap
 - Alkaline batteries
 - fertilizers
 - None
- Source of malic acid is:
 - apples
 - Urine
 - Fats
 - Butter
- Which is insoluble salt?
 - AgCl₂
 - NaNO₃
 - NaCl
 - Na₂SO₄
- What is the pOH of a 0.02 M Ca(OH)₂?
 - 1.397
 - 1.698
 - 12.31
 - 12.61
- Acetic acid is used for:
 - Flavorings food
 - etching designs
 - making explosives
 - Cleaning metals

Q. No. 2. Give Answers of the following short questions:**5x2=10**

- Write the uses of magnesium hydroxide.
- Write the physical properties of bases.
- Why a salt is neutral, explain with example?
- Why H⁺ ion act as lewis acid?
- Define pH. What is the pH of water?

Q. No. 3. Long Question:

- Define salt. Give the characteristic properties of salts. **05**

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Test: Chemistry	Group-	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q1. Answer the following Short Question.**5x1=05****1. Which one of the following is a substitution reaction?**

- (a) halogenations of alkynes (b) halogenations of alkenes
(c) halogenations of alkanes (d) bromination of alkenes

2. Which is the simplest alkane?

- (a) CH₄ (b) C₃H₈ (c) C₂H₂ (d) C₂H₄

3. Traces of acetylene are present in coal gas about

- (a) 0.06% (b) 0.08% (c) 1.1% (d) 90%

4. Which one of the following is more reactive?

- (a) methane (b) ethane (c) ethene (d) acetylene

5. Alkanes do not react in

- (a) diffused sunlight (b) bright sunlight (c) dark (d) none of these

Q. No. 2. Give Answers of the following short questions:**5x2=10**

- Give the physical properties of alkanes.**
- How can you identify ethane from ethene?**
- Define unsaturated hydrocarbons? Give example.**
- What are alkenes? Give its general formula?**
- How alkyl halides are reduced?**

Q. No. 3. Long Question:

Write the uses of alkenes.

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Test: Chemistry	Group-	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q1. Answer the following Short Question.**5x1=05****1. Which one of the following is more reactive?**

- (a) methane (b) ethane (c) ethene (d) acetylene

2. Traces of acetylene are present in coal gas about

- (a) 0.06% (b) 0.08% (c) 1.1% (d) 90%

3. Which is the simplest alkane?

- (a) CH₄ (b) C₃H₈ (c) C₂H₂ (d) C₂H₄

4. Alkanes do not react in

- (a) diffused sunlight (b) bright sunlight (c) dark (d) none of these

5. Which one of the following is a substitution reaction?

- (a) halogenations of alkynes (b) halogenations of alkenes
(c) halogenations of alkanes (d) bromination of alkenes

Q. No. 2. Give Answers of the following short questions:**5x2=10**

- What are alkenes? Give its general formula?**
- How alkyl halides are reduced?**
- Define unsaturated hydrocarbons? Give example.**
- Give the physical properties of alkanes.**
- How can you identify ethane from ethene?**

Q. No. 3. Long Question:

Write the uses of alkenes.

05

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 05**

- Bases have taste.
 - Sour
 - Bitter
 - Foul
 - Pleasant
- Which one of the following is not acid?
 - NH₃
 - HCl
 - CH₃COOH
 - H₂SO₄
- All bases turn red litmus.
 - Red
 - White
 - Blue
 - Pink
- Which one of the following is Lewis base?
 - NH₃
 - BF₃
 - H⁺
 - AlCl₃
- Who proved that presence of hydrogen as the main constituent of all acid?
 - Lavoisier
 - H. Davy
 - Dalton
 - Lewis

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why H⁺ ion acts as Lewis acid?
- Give characteristics properties of acids.
- Write the limitations of Arrhenius concept.
- What is adduct? Give example.
- Define amphoteric compounds. Give an example.

Q. No.3. Explain the Lewis concept of acids and bases.**05****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 05**

- All bases turn red litmus.
 - Red
 - White
 - Blue
 - Pink
- Who proved that presence of hydrogen as the main constituent of all acid?
 - Lavoisier
 - H. Davy
 - Dalton
 - Lewis
- Bases have taste.
 - Sour
 - Bitter
 - Foul
 - Pleasant
- Which one of the following is not acid?
 - NH₃
 - HCl
 - CH₃COOH
 - H₂SO₄
- Which one of the following is Lewis base?
 - NH₃
 - BF₃
 - H⁺
 - AlCl₃

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define acid according to H. Davy.
- Give the characteristics of bases.
- What is conjugate acid and base? Give examples.
- Prove that water is amphoteric compound.
- Why BF₃ behaves as a Lewis acid?

Q. No.3. What is acid and base according to Arrhenius? Explain with examples.**05**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 05**

- SO₂ gas used to manufacturing of
 - H₂SO₄
 - HNO₃
 - HCl
 - H₂CO₃
- The value of K_c depends only on
 - pressure
 - temperature
 - atmosphere
 - none of these
- The reaction which have comparable amount of reactant and products at equilibrium state have
 - Very small K_c value
 - moderate K_c value
 - very long K_c value
 - none of these
- For the reaction of PCl₃ and Cl₂ to form PCl₅ the unit of K_c are
 - mol dm⁻³
 - mol⁻¹ dm⁻³
 - mol⁻¹ dm³
 - mol dm³
- When a reaction ceases to proceed further, it is called
 - Chemical state
 - static state
 - physical state
 - none of these

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Give the characteristics of reversible reaction.
- Why at equilibrium state reaction does not stop?
- How direction of reaction can be predicted?
- Write the equilibrium constant expression for following equations
 - $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$
 - $\text{CO} + 3\text{H}_2\text{O} \rightleftharpoons \text{CH}_4 + \text{H}_2\text{O}$
- What do you mean by equilibrium constant?

Q. No.3. a) State the law of mass action and derive the expression for equilibrium constant for a general reaction.**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 05**

- SO₂ gas used to manufacturing of
 - H₂SO₄
 - HNO₃
 - HCl
 - H₂CO₃
- The value of K_c depends only on
 - pressure
 - temperature
 - atmosphere
 - none of these
- The reaction which have comparable amount of reactant and products at equilibrium state have
 - Very small K_c value
 - moderate K_c value
 - very long K_c value
 - none of these
- For the reaction of PCl₃ and Cl₂ to form PCl₅ the unit of K_c are
 - mol dm⁻³
 - mol⁻¹ dm⁻³
 - mol⁻¹ dm³
 - mol dm³
- The oxidation of carbon monoxide goes to completion at 1000 K the value of K_c is
 - 3.0 x 10⁻⁹
 - 2.2 x 10²²
 - 1.8 x 10⁻¹⁹
 - 1.1 x 10²⁴

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define irreversible reaction.
- How direction of reaction can be predicted?
- Write the equilibrium constant expression for following equations
 - $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$
 - $\text{CO} + 3\text{H}_2\text{O} \rightleftharpoons \text{CH}_4 + \text{H}_2\text{O}$
- If the reaction has larger value of K_c will it go to completion and why?
- What is static equilibrium? Explain with example.

Q. No.3. a) Write the macroscopic characteristics of dynamic equilibrium.**5**

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Test: Chemistry Time: 40 Min. Name: _____	Group Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Normally ran water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas
- Infrared radiation emitted by the earth are absorbed by
a) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Earth has natural systems.
a) one b) two c) three d) four

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Why CO₂ gas is called green house gas?
- How aquatic life is affected by acid rain
- Write the composition of dry air.
- Point out the two serious effects of ozone depletion.
- What is incineration?
- State the major source of CO and CO₂.
- Explain the phenomenon of decreasing temperature in troposphere.
- Define atmosphere.

Q. No.3.

- How acid rain forms? What are its effects? 5
- How ozone layer is depleted? 4

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Test: Chemistry Time: 40 Min. Name: _____	Group Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Earth has natural systems.
a) one b) two c) three d) four
- Infrared radiation emitted by the earth are absorbed by
b) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Normally ran water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 6 = 16**

- Name the major constituents of troposphere.
- What do you mean by air pollutant?
- Why flood risk increases?
- CO₂ is hidden enemy, Explain.
- What are the threats are to human health to SO₂ gas as air pollutant?
- Hoe acid rain increases the acidity of soil?
- What do you mean by ozone hole?
- How ozone layer is formed?

Q. No.3.

- How acid rain forms? What are its effects? 5
- How ozone layer is depleted? 4

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Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Normally rain water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas
- Infrared radiation emitted by the earth are absorbed by
a) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Earth has natural systems.
a) one b) two c) three d) four

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Why CO₂ gas is called green house gas?
- How aquatic life is affected by acid rain?
- Write the composition of dry air.
- Point out the two serious effects of ozone depletion.
- What is incineration?
- State the major source of CO and CO₂.
- Explain the phenomenon of decreasing temperature in troposphere.
- Define atmosphere & environment.

Q. No.3.

- How acid rain forms? What are its effects? 5
- Write the procedure to identify carboxylic acid using sodium carbonate test. 4

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Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Earth has natural systems.
a) one b) two c) three d) four
- Infrared radiation emitted by the earth are absorbed by
a) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Normally rain water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Name the major constituents of troposphere.
- What do you mean by air pollutant?
- Why flood risk increases?
- CO₂ is hidden enemy, Explain.
- What are the threats are to human health to SO₂ gas as air pollutant?
- How acid rain increases the acidity of soil?
- What do you mean by ozone hole?
- How ozone layer is formed?

Q. No.3.

- How ozone layer is depleted? 5
- Write the procedure to identify phenol using ferric chloride test. 4

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Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Normally rain water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas
- Infrared radiation emitted by the earth are absorbed by
a) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Earth has natural systems.
a) one b) two c) three d) four

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Why CO₂ gas is called green house gas?
- How aquatic life is affected by acid rain?
- Write the composition of dry air.
- Point out the two serious effects of ozone depletion.
- What is incineration?
- State the major source of CO and CO₂.
- Explain the phenomenon of decreasing temperature in troposphere.
- Define atmosphere & environment.

Q. No.3.

- How acid rain forms? What are its effects? 5
- Write the procedure to identify carboxylic acid using sodium carbonate test. 4

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Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Earth has natural systems.
a) one b) two c) three d) four
- Infrared radiation emitted by the earth are absorbed by
a) CO₂ and H₂O b) N₂ and O₂ c) O₂ and CO₂ d) CO₂ and N₂
- Which of the following is secondary pollutant?
a) SO₃ b) CH₄ c) NO d) None of these
- High concentration of which metal clogs fish gills.
a) Zinc b) aluminum c) sodium d) Copper
- Normally rain water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Name the major constituents of troposphere.
- What do you mean by air pollutant?
- Why flood risk increases?
- CO₂ is hidden enemy, Explain.
- What are the threats are to human health to SO₂ gas as air pollutant?
- How acid rain increases the acidity of soil?
- What do you mean by ozone hole?
- How ozone layer is formed?

Q. No.3.

- How ozone layer is depleted? 5
- Write the procedure to identify phenol using ferric chloride test. 4

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Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- The formula of phosphoric acid is.
 - H_2PO_4
 - H_3PO_4
 - H_2SO_3
 - CH_3COOH
- Which of the following will use Natural food preservatives are:
 - Salts
 - Sugar
 - Alcohols
 - all of these
- Which one of the following salt is used in manufacture of flint glass?
 - CaO
 - $CaCl_2$
 - $NaClO_3$
 - KNO_3
- Which of the following is neutral salt?
 - $NaHPO_4$
 - $Zn(OH)NO_3$
 - KCl
 - none of these
- Formula of potash alum is
 - $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24 H_2O$
 - $KClO_3$
 - $CuSO_4 \cdot 5H_2O$
 - none of these
- Example of mixed salt
 - $Pb(CH_3COO)_2$
 - $Ca(OCl)Cl$
 - NH_4NO_3
 - $KHSO_4$
- Which one is insoluble salt?
 - NaCl
 - $CuSO_4$
 - $CuCO_3$
 - Na_2CO_3

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define pH. What is pH of water?
- How insoluble salts are prepared?
- Why H^+ ion acts as a Lewis acid?
- How the basic salts turn into normal salts?
- Give the limitations of Arrhenius concept.
- Give the uses of calcium oxide.
- What are complex salts?

Q. No.3. a) Write the procedure to standardize the give HCl solution volumetrically.**4****b) Define acid and base according to Lowry-Bronsted concept. Explain with examples. 5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Which acid is used for etching designs on copper plates?
 - H_2SO_4
 - HNO_3
 - HCl
 - CH_3COOH
- You want to dry a gas which of the following will use Natural food preservatives are:
 - Salts
 - Sugar
 - Alcohols
 - all of these
- When acid react with carbonates and bicarbonates which gas is evolved?
 - H_2
 - CO_2
 - Cl_2
 - N_2
- Which one of the following salt is used in manufacture of flint glass?
 - CaO
 - $CaCl_2$
 - $NaClO_3$
 - KNO_3
- Example of mixed salt
 - $Pb(CH_3COO)_2$
 - $Ca(OCl)Cl$
 - NH_4NO_3
 - $KHSO_4$
- Which one is insoluble salt?
 - NaCl
 - $CuSO_4$
 - $CuCO_3$
 - Na_2CO_3
- The formula of phosphoric acid is
 - H_2PO_4
 - H_3PO_4
 - H_2SO_3
 - CH_3COOH

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- How insoluble salts are prepared?
- Give few characteristics of salts.
- Explain neutralization reaction according to Lewis concept.
- What are complex salts?
- How double salts are prepared?
- Prove that water is amphoteric compound.
- Why BF_3 behave as a Lewis acid?

Q. No.3. a) What is auto ionization? How it is used to establish the pH of water?**5****b) Write the observations to determine the exact molarity of Na_2CO_3 solution volumetrically. 4**

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Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- The formula of phosphoric acid is
a) H_2PO_4 b) H_3PO_4 c) H_2SO_3 d) CH_3COOH
- You want to dry a gas which of the following will use Natural food preservatives are:
a) Salts b) Sugar c) Alcohols d) all of these
- Which one of the following salt is used in manufacture of flint glass?
a) CaO b) $CaCl_2$ c) $NaClO_3$ d) KNO_3
- Which of the following is neutral salt?
a) $NaHPO_4$ b) $Zn(OH)NO_3$ c) KCl d) none of these
- Formula of potash alum is
a) $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24 H_2O$ b) $KClO_3$
c) $CuSO_4 \cdot 5H_2O$ d) none of these
- Example of mixed salt
a) $Pb(CH_3COO)_2$ b) $Ca(OCl)Cl$ c) NH_4NO_3 d) $KHSO_4$
- Which one is insoluble salt?
a) $NaCl$ b) $CuSO_4$ c) $CuCO_3$ d) Na_2CO_3

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define pH. What is pH of water?
- How insoluble salts are prepared?
- Why H^+ ion acts as a Lewis acid?
- How the basic salts turn into normal salts?
- Give the limitations of Arrhenius concept.
- Give the uses of calcium oxide.
- What are complex salts?

Q. No.3. a) Write the procedure to standardize the give HCl solution volumetrically.**5****b) Define acid and base according to Lowry-Bronsted concept. Explain with examples. 4****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Which acid is used for etching designs on copper plates.
a) H_2SO_4 b) HNO_3 c) HCl d) CH_3COOH
- You want to dry a gas which of the following will use Natural food preservatives are:
a) Salts b) Sugar c) Alcohols d) all of these
- When acid react with carbonates and bicarbonates which gas is evolved?
a) H_2 b) CO_2 c) Cl_2 d) N_2
- Which one of the following salt is used in manufacture of flint glass?
a) CaO b) $CaCl_2$ c) $NaClO_3$ d) KNO_3
- Example of mixed salt
a) $Pb(CH_3COO)_2$ b) $Ca(OCl)Cl$ c) NH_4NO_3 d) $KHSO_4$
- Which one is insoluble salt?
a) $NaCl$ b) $CuSO_4$ c) $CuCO_3$ d) Na_2CO_3
- The formula of phosphoric acid is
a) H_2PO_4 b) H_3PO_4 c) H_2SO_3 d) CH_3COOH

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- How insoluble salts are prepared?
- Give few characteristics of salts.
- Explain neutralization reaction according to Lewis concept.
- What are complex salts?
- How double salts are prepared?
- Prove that water is amphoteric compound.
- Why BF_3 behave as a Lewis acid?

Q. No.3. a) What is auto ionization? How it is used to establish the pH of water?**5****b) Write the observations to determine the exact molarity of Na_2CO_3 solution volumetrically. 4**

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Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Atmospheric water contain about world water
a) 0.2 % b) 0.001 % c) 97 % d) 2.1 %
- Which of the following salt causes temporary hardness
a) CaCl_2 b) MgSO_4 c) MgCl_2 d) none of these
- Acute cadmium poisoning causes
a) gastro b) cancer c) kidney damage d) none
- Rapid growth of algae in water bodies is because of detergents having?
a) carbonate salt b) sulphonic acid salts c) Sulphate salts d) Phosphate salt
- Which one of the following diseases causes liver inflammation?
a) typhoid b) jaundice c) Cholera d) hepatitis

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- What is scum?
- How detergents make the water unfit for aquatic life?
- How lime stone dissolves in water?
- What is the function of fertilizers?
- What is dysentery?
- Explain the chemistry of removing the temporary hardness by boiling water?
- Define fluorosis.
- How water rises in plants?

Q. No.3. a). Explain the methods removing permanent hardness?**4****b). What is the chemistry of swimming pool cleanliness?****3****ATP****How can you identify the saturated and unsaturated organic compounds by KMnO_4 test? 2****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- A disease that causes bone and tooth damage is
a) cholera b) fluorosis c) hepatitis d) jaundice
- Water dissolves non-ionic compounds by
a) Ion-ion forces b) ion-dipole forces c) dipole-dipole forces d) hydrogen bonding
- Formula of slaked lime is
a) CaC_2 b) Ca(OH)_2 c) Mg(OH)_2 d) none of these
- Mercury poisoning causes
a) High blood pressure b) neurological damages c) Kidney damage d) None of these
- Hepatitis A and E can be transmitted by
a) fluorides b) contaminated water c) protozoa d) none of these

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 8 = 16**

- Why non polar compounds are insoluble in water?
- Why pesticides are used?
- What are the main causes of hardness in water?
- Mention the disadvantages of detergents.
- How addition of Na_2CO_3 removes permanent hardness of water?
- Why water molecule is polar?
- What do you mean by boiler scales?
- What are industrial effluents?

Q. No.3. a). Explain the reasons, water are considered a universal solvent.**4****b). Write the disadvantages of hard water.****3****ATP****How can you demonstrate that sugar decomposes into elements or other compounds? 2**

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Test: Chemistry Time: 1 Hour Name: _____	Class: 10th Max. Marks: 30 Section: _____
Date: _____	

Q. No.1. Choose the Correct Answer.

1 x 7 = 7

1. When a reaction ceases to proceed further, it is called
 a) Chemical state b) static state c) physical state d) none of these
2. The % age of nitrogen and oxygen in our atmosphere is
 a) 98% b) 90% c) 85% d) 99%
3. The oxidation of carbon monoxide goes to completion at 1000 K the value of K_c is
 a) 3.0×10^{-9} b) 2.2×10^{22} c) 1.8×10^{-19} d) 1.1×10^{24}
4. For the reaction of PCl₃ and Cl₂ to form PCl₅ the unit of K_c are
 a) mol dm⁻³ b) mol⁻¹ dm⁻³ c) mol⁻¹ dm³ d) mol dm³
5. The reaction which have comparable amount of reactant and products at equilibrium state have
 a) Very small K_c value b) moderate K_c value
 c) very long K_c value d) none of these
6. SO₂ gas used to manufacturing of
 a) H₂SO₄ b) HNO₃ c) HCl d) H₂CO₃
7. The value of K_c depends only on
 a) pressure b) temperature c) atmosphere d) none of these

Q. No.2. Give Short Answers of the following questions.

2 x 7 = 14

1. Give the characteristics of reversible reaction.
2. Why at equilibrium state reaction does not stop?
3. How direct of reaction can be predicted?
4. Write the equilibrium constant expression for following equations
 a) $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$ b) $\text{CO} + 3\text{H}_2\text{O} \rightleftharpoons \text{CH}_4 + \text{H}_2\text{O}$
5. If the reaction has larger value of K_c will it go to completion and why?
6. What do you mean by equilibrium constant?
7. What is static equilibrium? Explain with example.

Q. No.3. a) State the law of mass action and derive the expression for equilibrium constant for a general reaction.

5

b) Write the observations.

4

Metallic ion	Sodium ion	Calcium ion	Barium ion	Copper ion
Colour imparted to the flame				

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Test: Chemistry	Class: 10th
Time: 1 Hour	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.

1 x 7 = 7

1. SO₂ gas used to manufacturing of
a) H₂SO₄ b) HNO₃ c) HCl d) H₂CO₃
2. The value of K_c depends only on
a) pressure b) temperature c) atmosphere d) none of these
3. The reaction which have comparable amount of reactant and products at equilibrium state have
a) Very small K_c value b) moderate K_c value
c) very long K_c value d) none of these
4. For the reaction of PCl₃ and Cl₂ to form PCl₅ the unit of K_c are
a) mol dm⁻³ b) mol⁻¹ dm⁻³ c) mol⁻¹ dm³ d) mol dm³
5. The oxidation of carbon monoxide goes to completion at 1000 K the value of K_c is
a) 3.0 x 10⁻⁹ b) 2.2 x 10²² c) 1.8 x 10⁻¹⁹ d) 1.1 x 10²⁴
6. When a reaction ceases to proceed further, it is called
a) Chemical state b) static state c) physical state d) none of these
7. The % age of nitrogen and oxygen in our atmosphere is
a) 98% b) 90% c) 85% d) 99%

Q. No.2. Give Short Answers of the following questions.

2 x 7 = 14

1. Give the characteristics of reversible reaction.
2. Why at equilibrium state reaction does not stop?
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4. Write the equilibrium constant expression for following equations
b) $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$ b) $\text{CO} + 3\text{H}_2\text{O} \rightleftharpoons \text{CH}_4 + \text{H}_2\text{O}$
5. If the reaction has larger value of K_c will it go to completion and why?
6. What do you mean by equilibrium constant?
7. What is static equilibrium? Explain with example.

Q. No.3. a) Write the macroscopic characteristics of dynamic equilibrium.

5

b) Write the procedure to standardize the given NaOH solution volumetrically.

4

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Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th A Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- About 99% atmosphere's mass lies within:
 - 30 Km
 - 35 Km
 - 11Km
 - 15Km
- Accumulation of CO₂ gas atmospheric temperature increasing every year.
 - 0.5 °C
 - 0.05°C
 - 5°C
 - none
- Which of the following is poisonous gas
 - O₂
 - O₃
 - CO
 - N₂
- Atmospheric gases absorbed sunlight
 - 26%
 - 32%
 - 18%
 - 6%
- % age of Argon in atmosphere is
 - 0.03
 - 0.93
 - 20.9
 - 78

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How the temperature of atmosphere is maintained?
- Why CO is considered a health hazard?
- What are the characteristics of stratosphere?
- Why increasing concentration of CO₂ is alarming for us?
- Write the effects of global warning?

Q. No.3. Write down the significance of atmospheric gases**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of these pollutants are found in car exhaust fumes?
 - CO
 - O₃
 - NO₂
 - SO₂
- Accumulation of CO₂ gas atmospheric temperature increasing every year.
 - 0.5 °C
 - 0.05°C
 - 5°C
 - none
- Which of the following is poisonous gas
 - O₂
 - O₃
 - CO
 - N₂
- Atmospheric gases absorbed sunlight
 - 26%
 - 32%
 - 18%
 - 6%
- % age of Argon in atmosphere is
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(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How the temperature of atmosphere is maintained?
- Why CO is considered a health hazard?
- What are the characteristics of stratosphere?
- Why increasing concentration of CO₂ is alarming for us?
- Name the major constituents of troposphere.

Q. No.3. What is greenhouse effect and global warning?**5**

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Test: Chemistry Time: 40 Min. Name: _____	G-1 Date: _____	Class: 10 th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Potable water on earth is only.
 - 0.2%
 - 0.1%
 - 0.3%
 - 0.01%
- The maximum density of water at.
 - 4°C
 - 3°C
 - 100°C
 - 101°C
- Temporary hardness is because of.
 - Ca(HCO₃)₂
 - CaCO₃
 - MgCO₃
 - MgSO₄
- Permanent hardness is removed by adding.
 - Na₂-Zeolite
 - Soda lime
 - Lime water
 - Quick lime
- Benzene is not dissolve in.
 - Ether
 - Octane
 - Alcohol
 - Water

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why non ionic polar compounds are soluble in water?
- How detergents make water unfit for aquatic life?
- How lime stone dissolve in water?
- What are the causes of hardness in water?
- Differentiate btw soft and hard water.

Q. No.3. Explain the methods of removing permanent hardness.**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	G-2 Date: _____	Class: 10 th A Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Benzene is not dissolve in.
 - Ether
 - Octane
 - Alcohol
 - Water
- Potable water on earth is only.
 - 0.2%
 - 0.1%
 - 0.3%
 - 0.01%
- The maximum density of water at.
 - 4°C
 - 3°C
 - 100°C
 - 101°C
- Permanent hardness is removed by adding.
 - Na₂-Zeolite
 - Soda lime
 - Lime water
 - Quick lime
- Temporary hardness is because of.
 - Ca(HCO₃)₂
 - CaCO₃
 - MgCO₃
 - MgSO₄

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What are the causes of hardness in water?
- How hard water hampers the cleaning action of soap?
- Differentiate btw soft and hard water.
- How lime stone dissolve in water?
- What is the function of fertilizers?

Q. No.3. Write note on water born diseases.**5**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Pitch is black residue of
 - Coke
 - Coal tar
 - coal
 - coal gas
- The functional group C-O-C found in
 - Esters
 - aldehydes
 - carboxylic acid
 - alcohol
- Carbon present in lignite is
 - 60%
 - 70%
 - 80%
 - 90%
- Which of the following is aromatic compound?
 - Benzene
 - phenol
 - toluene
 - all of these
- Which of the following is the source of proteins?
 - coal
 - butter
 - barley
 - pulses
- How much energy is required to break C-O bond.
 - 355 KJ mol⁻¹
 - 200 KJ mol⁻¹
 - 452 KJ mol⁻¹
 - 351 KJ mol⁻¹
- Each person's sweat contains a unique blend of
 - Tartaric acid
 - H₂SO₄
 - Carboxylic acid
 - none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define catenation and isomerism.
- What is ester linkage?
- What is the importance of coal tar?
- How alkyl radical are formed? Give its general formula.
- What are aliphatic compounds?
- What is the difference b/w aldehyde and ketones?
- Define functional group. Give example.

Q. No.3. a) Write down the characteristics of homologous series.**5****b) How can you determine the exact molarity of oxalic acid volumetrically.****4****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 30
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Pitch is black residue of
 - Coke
 - Coal tar
 - coal
 - coal gas
- The functional group C-O-C found in
 - Esters
 - aldehydes
 - carboxylic acid
 - alcohol
- Carbon present in lignite is
 - 60%
 - 70%
 - 80%
 - 90%
- Which of the following is aromatic compound?
 - Benzene
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 - all of these
- Which of the following is the source of proteins?
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 - 355 KJ mol⁻¹
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- Each person's sweat contains a unique blend of
 - Tartaric acid
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(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define catenation and isomerism.
- What is vital force theory?
- What is the importance of coal tar?
- How alkyl radical are formed? Give its general formula.
- What is ester group? Write the formula of ethyle acetate.
- What are heterocyclic compound? Give example.
- Define functional group. Give example.

Q. No.3. a) Write are the general properties of organic compounds?**5****b) How can you demonstrate that some natural substances are weak acids.****4**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- About 99% atmosphere's mass lies within:
 - 30 Km
 - 35 Km
 - 11Km
 - 15Km
- Accumulation of CO₂ gas atmospheric temperature increasing every year.
 - 0.5 °C
 - 0.05°C
 - 5°C
 - none
- Which of the following is poisonous gas
 - O₂
 - O₃
 - CO
 - N₂
- Atmospheric gases absorbed sunlight
 - 26%
 - 32%
 - 18%
 - 6%
- % age of Argon in atmosphere is
 - 0.03
 - 0.93
 - 20.9
 - 78

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How the temperature of atmosphere is maintained?
- Why CO is considered a health hazard?
- What are the characteristics of stratosphere?
- Why increasing concentration of CO₂ is alarming for us?
- Write the effects of global warning?

Q. No.3. Write down the significance of atmospheric gases**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th B
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of these pollutants are found in car exhaust fumes?
 - CO
 - O₃
 - NO₂
 - SO₂
- Accumulation of CO₂ gas atmospheric temperature increasing every year.
 - 0.5 °C
 - 0.05°C
 - 5°C
 - none
- Which of the following is poisonous gas
 - O₂
 - O₃
 - CO
 - N₂
- Atmospheric gases absorbed sunlight
 - 26%
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- % age of Argon in atmosphere is
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(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How the temperature of atmosphere is maintained?
- Why CO is considered a health hazard?
- What are the characteristics of stratosphere?
- Why increasing concentration of CO₂ is alarming for us?
- Name the major constituents of troposphere.

Q. No.3. What is greenhouse effect and global warning?**5**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: B

Q. No. 1. Choose the Correct Answer.**1 x 5 = 5**

- At initial stage the rate of forward reaction is:
 - very slow
 - very fast
 - negligible
 - none
- The major components of atmosphere:
 - CO₂
 - nitrogen & oxygen
 - oxygen
 - none
- Which natural process is responsible for existence of life of earth
 - exchange of gases
 - burning
 - both a & b
 - none of these
- For a reaction between PCI₃ and Cl₂ to form PCI₅, the units of K_c are:
 - mol dm⁻³
 - mol⁻¹ dm⁻³
 - mol⁻¹ dm³
 - mol dm³
- The reactions in which the products do not recombine to form reactants are called:
 - reversible
 - irreversible
 - forward
 - reverse

(SUBJECTIVE TYPE)**Q. No. 2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define Irreversible reaction. Give its characteristics.
- Write the equilibrium constant expression for following reaction

$$\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$$
- Define chemical equilibrium state.
- Why at equilibrium state reaction does not stop?
- How a dynamic equilibrium is established?

Q. No. 3. State the law of Mass Action and drive the expression for equilibrium constant for a general reaction 05

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: B

Q. No. 1. Choose the Correct Answer.**1 x 5 = 5**

- At initial stage the rate of forward reaction is:
 - very slow
 - very fast
 - negligible
 - none
- The major components of atmosphere:
 - CO₂
 - nitrogen & oxygen
 - oxygen
 - none
- Which natural process is responsible for existence of life of earth
 - exchange of gases
 - burning
 - both a & b
 - none of these
- For a reaction between PCI₃ and Cl₂ to form PCI₅, the units of K_c are:
 - mol dm⁻³
 - mol⁻¹ dm⁻³
 - mol⁻¹ dm³
 - mol dm³
- The reactions in which the products do not recombine to form reactants are called:
 - reversible
 - irreversible
 - forward
 - reverse

(SUBJECTIVE TYPE)**Q. No. 2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define reversible reaction. Give its characteristics.
- Write the equilibrium constant expression for following reaction

$$\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$$
- Define chemical equilibrium state.
- Why at equilibrium state reaction does not stop?
- How direction of reaction can be predicted?

Q. No. 3. How extent of reaction is predicted by equilibrium constant.

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Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- Which is proteins
a) Pulses b) beans c) butter d) both a and b
- Which into is absorbed directly small intestine into blood stream?
a) Starch b) lipide c) fats d) glucose
- Smell of methyl butanoate is like
a) apple b) pineapple c) orange d) both a and b
- Deficiency of vitamin D causes.
a) Night blindness b) Rickets c) Cancer d) none of these
- Which of the following is reducing sugar?
a) glucose b) maltose c) fructose d) starch
- When glucose and fructose combine they produce
a) Starch b) cellulose c) vitamin d) none of these
- The organic compound used to control bleeding are
a) Vitamin b) proteins c) lipids d) glycerides

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Define lipids. Give its general formula.
- How margarine is formed?
- Which elements are found in proteins?
- Why RNA are called messenger?
- How protein is formed?
- What is the significance of vitamins?
- How gelatin is obtained?

- Q. No.3. a) Write the commercial used of microbial and lactase enzymes. 5**
b) How can you identify aldehydes using Fehling solution? 4

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Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 30 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 7 = 7**

- The most important oligosaccharide is
a) glucose b) Fructose c) maltose d) sucrose
- Which one is triglyceride?
a) Portions b) lipids c) vitamins d) DNA
- Which of the following vitamins is water soluble?
a) vitamins A b) vitamins B c) vitamins C d) vitamins D
- Smell of ethyl butanoate is like
a) apple b) pine apple c) orange d) foul
- Which of the following is high energy food
a) lipids b) proteins c) carbohydrates d) none
- Carbohydrates provide energy per gram.
a) 17 KJ b) 170 KJ c) 200 KJ d) 1700 KJ
- Formula of palmitic acid is
a) $C_{12}H_{22}O_{11}$ b) $C_{15}H_{31}COOH$ c) $(NH_2)_2CO$ d) $C_{17}H_{35}COOH$

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 7 = 14**

- Give the characteristics of monosaccharide's.
- What is the function of DNA?
- What is difference between glucose and fructose?
- What is the importance and source of vitamin A?
- How proteins are formed?
- What are the components of nucleotide?
- Write the used of amylase enzymes.

- Q. No.3. a) Write the sources and uses of proteins? 5**

- b) How can you identify ketones using 2-4 dinitrophenyl hydrazine test? 4**

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Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th B Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- About 99% atmosphere's mass lies within:
 - 30 Km
 - 35 Km
 - 11Km
 - 15Km
- Accumulation of CO₂ gas atmospheric temperature increasing every year.
 - 0.5 °C
 - 0.05°C
 - 5°C
 - none
- Which of the following is poisonous gas
 - O₂
 - O₃
 - CO
 - N₂
- Atmospheric gases absorbed sunlight
 - 26%
 - 32%
 - 18%
 - 6%
- % age of Argon in atmosphere is
 - 0.03
 - 0.93
 - 20.9
 - 78

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How the temperature of atmosphere is maintained?
- Why CO is considered a health hazard?
- What are the characteristics of stratosphere?
- Why increasing concentration of CO₂ is alarming for us?
- Write the effects of global warning?

Q. No.3. Write down the significance of atmospheric gases**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th B Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of these pollutants are found in car exhaust fumes?
 - CO
 - O₃
 - NO₂
 - SO₂
- Accumulation of CO₂ gas atmospheric temperature increasing every year.
 - 0.5 °C
 - 0.05°C
 - 5°C
 - none
- Which of the following is poisonous gas
 - O₂
 - O₃
 - CO
 - N₂
- Atmospheric gases absorbed sunlight
 - 26%
 - 32%
 - 18%
 - 6%
- % age of Argon in atmosphere is
 - 0.03
 - 0.93
 - 20.9
 - 78

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How the temperature of atmosphere is maintained?
- Why CO is considered a health hazard?
- What are the characteristics of stratosphere?
- Why increasing concentration of CO₂ is alarming for us?
- Name the major constituents of troposphere.

Q. No.3. What is greenhouse effect and global warning?**5**

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Test: Chemistry	G-1	Class: 10th B
Time: 40 Min.		Max. Marks: 20
Name: _____	Date: _____	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Potable water on earth is only.
 - 0.2%
 - 0.1%
 - 0.3%
 - 0.01%
- The maximum density of water at.
 - 4°C
 - 3°C
 - 100°C
 - 101°C
- Temporary hardness is because of.
 - Ca(HCO₃)₂
 - CaCO₃
 - MgCO₃
 - MgSO₄
- Permanent hardness is removed by adding.
 - Na₂-Zeolite
 - Soda lime
 - Lime water
 - Quick lime
- Benzene is not dissolve in.
 - Ether
 - Octane
 - Alcohol
 - Water

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why non ionic polar compounds are soluble in water?
- How detergents make water unfit for aquatic life?
- How lime stone dissolve in water?
- What are the causes of hardness in water?
- Differentiate btw soft and hard water.

Q. No.3. Explain the methods of removing permanent hardness.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	G-2	Class: 10th B
Time: 40 Min.		Max. Marks: 20
Name: _____	Date: _____	Section: _____

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Benzene is not dissolve in.
 - Ether
 - Octane
 - Alcohol
 - Water
- Potable water on earth is only.
 - 0.2%
 - 0.1%
 - 0.3%
 - 0.01%
- The maximum density of water at.
 - 4°C
 - 3°C
 - 100°C
 - 101°C
- Permanent hardness is removed by adding.
 - Na₂-Zeolite
 - Soda lime
 - Lime water
 - Quick lime
- Temporary hardness is because of.
 - Ca(HCO₃)₂
 - CaCO₃
 - MgCO₃
 - MgSO₄

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What are the causes of hardness in water?
- How hard water hampers the cleaning action of soap?
- Differentiate btw soft and hard water.
- How lime stone dissolve in water?
- What is the function of fertilizers?

Q. No.3. Write note on water born diseases.**5**

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Test: Chemistry Time: 40 Min. Name: _____	Group-A Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Normally rain water is weakly acidic because of
a) SO₃ gas b) CO₂ gas c) NO₂ gas d) SO₂ gas
- High conc of which metal clogs the fish gills
a) Aluminum b) Chromium c) Radium d) Hg
- Acid rain consist of H₂SO₄ and HNO₃ reduce the pH of rain water to
a) 5.6 b) 6 c) 4 d) 3
- Which gas has irritating smell?
a) CO₂ b) SO₂ c) NO₂ d) O₂
- Which gas is called life gas for plants?
a) O₂ gas b) N₂ gas c) CO₂ gas d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How aquatic life is affected by acid rain?
- What do you mean by ozone hole?
- How ozone layer is being depleted by chlorofluoro carbon?
- Write the two serious effects of ozone depletion.
- What is incineration?

Q. No.3. Define acid rain. What are its effects?**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Group-B Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- 80% of SO₂ gas is released by the combustion of
a) coal b) petroleum products c) forest fire d) both a and b
- Which one of these pollutants are not found in car exhaust fumes?
a) CO b) O₃ c) NO₃ d) SO₂
- Acid rain affects the aquatic life by clogging fish gills because of
a) Lead metal b) Aluminum metal c) copper metal d) none
- Buildings are being damaged by acid rain because it attack.
a) CaSO₄ b) CaCO₃ c) CaCl₂ d) CaO
- The process by which atmospheric nitrogen is turned into nitrates in the soil is called
a) nitrification b) fixing c) oxidations d) reduction

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How acid rain is produced?
- How ozone layer is forms in stratosphere?
- What threats are to human health due to SO₂ gas as air pollutant?
- What is flue gas?
- Write two effects of acid rain.

Q. No.3. a) Write note on depletion of ozone layer. How we can prevent it?**5**

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Test: Chemistry Time: 40 Min. Name: _____	Group-A Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of the following is triglyceride?
a) Carbohydrates b) Portions c) lipids d) vitimins
- Which one of the following vitamins is water soluble?
a) Vitamin A b) Vitamin C c) Vitamin D d) none
- Deficiency of Vitamin E cause.
a) Scurvy b) rickets c) night blindness d) anemia
- Formula of stearic acid is
a) C₁₇ H₃₅COOH b) C₁₇H₃₇COOH c) CH₃COOH d) none
- Structure of DNA was discover by
a) S. William b) J. Watson c) Dalton d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is function of RNA?
- What is the significance of vitamin?
- Give uses and sources of animal fats.
- What are vegetable oils? How they are hydrogenated?
- What are the advantages of water soluble vitamins?

Q. No.3. a) What is the structure and function of DNA?**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Group-A Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which one of the following is triglyceride?
a) Carbohydrates b) Portions c) lipids d) vitimins
- Which one of the following vitamins is water soluble?
a) Vitamin A b) Vitamin C c) Vitamin D d) none
- Deficiency of Vitamin E cause.
a) Scurvy b) rickets c) night blindness d) anemia
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a) C₁₇ H₃₅COOH b) C₁₇H₃₇COOH c) CH₃COOH d) none
- Structure of DNA was discover by
a) S. William b) J. Watson c) Dalton d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is function of RNA?
- What is the significance of vitamin?
- Give uses and sources of animal fats.
- What are vegetable oils? How they are hydrogenated?
- What are the advantages of water soluble vitamins?

Q. No.3. a) What is the structure and function of DNA?**5**

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Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th B Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Night blindness is because of deficiency of
a) Vitamin A b) Vitamin B c) Vitamin C d) Vitamin D
- Which one of the following is a triglyceride?
a) Carbohydrates b) proteins c) lipids d) Vitamins
- Animal fats are used in
a) Medicine b) bakery items c) Soap industry d) none
- The comical formula of stearic acid is
a) $C_{15}H_{31}COOH$ b) $C_{15}H_{35}COOH$ c) $C_{17}H_{35}COOH$ d) none
- Rancid butter has a foul smell because of
a) Malic acid b) butanoic acid c) palmitic acid d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Plants are source of oils, Justify?
- What is the function of DNA?
- Give the general formula of lipids.
- Describe the source and uses of vitamin A.
- Why RNA is called a messenger?

Q. No.3. Give the importance of vitamins.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Date: _____	Class: 10th B Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Night blindness is because of deficiency of
a) Vitamin A b) Vitamin B c) Vitamin C d) Vitamin D
- Which one of the following is a triglyceride?
a) Carbohydrates b) proteins c) lipids d) Vitamins
- Animal fats are used in
a) Medicine b) bakery items c) Soap industry d) none
- The comical formula of stearic acid is
a) $C_{15}H_{31}COOH$ b) $C_{15}H_{35}COOH$ c) $C_{17}H_{35}COOH$ d) none
- Rancid butter has a foul smell because of
a) Malic acid b) butanoic acid c) palmitic acid d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How are protein formed?
- What is the significance of vitamins?
- Give the general formula of lipids.
- Write the use of amylase enzymes.
- Why RNA is called a messenger?

Q. No.3. Write note on DNA.**5**

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Test: Chemistry Time: 40 Min. Name: _____	Group B Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which is proteins
a) Pulses b) beans c) butter d) both a and b
- Which into is absorbed directly small intestine into blood stream?
a) Starch b) lipide c) fats d) glucose
- Smell of methyl butanoate is like
a) apple b) pineapple c) orange d) both a and b
- Deficiency of vitamin D causes.
a) Night blindness b) Rickets c) Cancer d) none of these
- Which of the following is reducing sugar?
a) glucose b) maltose c) fructose d) starch

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define lipids. Give its general formula.
- How margarine is formed?
- Which elements are found in proteins?
- Why RNA are called messenger?
- How protein is formed?

Q. No.3. a) Write note on oligosaccharides.**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry Time: 40 Min. Name: _____	Group A Date: _____	Class: 10th Max. Marks: 20 Section: _____
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Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- The most important oligosaccharide is
a) glucose b) Fructose c) maltose d) sucrose
- Which one is triglyceride?
a) Portions b) lipids c) vitamins d) DNA
- Which of the following vitamins is water soluble?
a) vitamins A b) vitamins B c) vitamins C d) vitamins D
- Smell of ethyl butanoate is like
a) apple b) pine apple c) orange d) foul
- Which of the following is high energy food
a) lipids b) proteins c) carbohydrates d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Give the characteristics of monosaccharide's.
- What is the function of DNA?
- What is difference between glucose and fructose?
- What is the importance and source of vitamin A?
- How proteins are formed?

Q. No.3. a) Write the sources and uses of proteins?**5**

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Test: Chemistry	Group-A	Class: 10th
Time: 40 Min.	Date: 09-07-2014	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 4 = 4**

- Which one of the following is saturated?
a) C₂H₄ b) C₃H₆ c) C₄H₈ d) C₅H₁₂
- Substitution reaction is the property of.
a) alkanes b) alkenes c) alkynes d) none
- Which of the following is used in drycleaning.
a) CCl₄ b) Chloroform c) acetone d) alcohol
- Methane form about 85% of.
a) Petroleum b) gobar gas c) natural gas d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define unsaturated hydrocarbons. Give examples.
- How alkyl halides are reduced?
- Write three physical properties of alkanes.
- Why alkanes are called paraffins?
- Why alkanes are used as fuel.

Q. No.3. What types of reactions are given by alkanes. Explain with reference to halogenations of alkanes.**06****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Group-B	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 4 = 4**

- Which of the following is used in drycleaning.
a) CCl₄ b) Chloroform c) acetone d) alcohol
- Methane form about 85% of.
a) Petroleum b) gobar gas c) natural gas d) none
- Which one of the following is saturated?
a) C₂H₄ b) C₃H₆ c) C₄H₈ d) C₅H₁₂
- Substitution reaction is the property of.
a) alkanes b) alkenes c) alkynes d) none

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why alkanes are used as fuel.
- Why alkanes are called paraffins?
- How alkyl halides are reduced?
- Define unsaturated hydrocarbons. Give examples.
- Write three physical properties of alkanes.

Q. No.3. What types of reactions are given by alkanes. Explain with reference to halogenations of alkanes.**06**

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Test: Chemistry	Group-A	Class: 10th A
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which of the following is hardest coal?
 - Lignite
 - peat
 - Anthracite
 - none
- Main components of natural gas is
 - Methane
 - Propane
 - butane
 - propane
- Which of the following does not contain protein
 - Pulses
 - potatoes
 - eggs
 - beans
- Pitch is the black residue of.
 - Coke
 - Coaltar
 - Coal
 - Coal gas
- Peat contains carbon
 - 60%
 - 70%
 - 80%
 - 90%

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Name the gases which are found in coal gas?
- What are aromatic compounds?
- Draw the structural formulae of n-butane and isobutene.
- Define heterocyclic compounds. Give examples.
- What is the importance of natural gas?

Q. No.3. How is coal formed? What are the different types of coal? Explain.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Group-B	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Coal gas is the mixture of.
 - CO and CH₄
 - CO, CH₄, CO₂
 - CO, H₂, CH₂
 - none
- Which of the following is synthetic fiber?
 - Cotton
 - nylon
 - wool
 - silk
- The percentage of carbon in lignite is.
 - 60%
 - 90%
 - 80%
 - None of these
- The Superior quality of coal is
 - Peat
 - lignite
 - Anthracite
 - none
- Coal Tar is the mixture of compounds more than.
 - 200
 - 300
 - 500
 - 100

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is ammoniacal liquor?
- How coal is formed?
- What are aromatic compounds? Give example.
- What is isomerism? Give the structure of isomers of butane.
- Which types of compounds are synthesized by plants?

Q. No.3. Write down the characteristics of organic compounds?**5**

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Test: Chemistry	Group-A	Class: 10th A
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which of the following is hardest coal?
 - Lignite
 - peat
 - Anthracite
 - none
- Main components of natural gas is
 - Methane
 - Propane
 - butane
 - propane
- Which of the following does not contain protein
 - Pulses
 - potatoes
 - eggs
 - beans
- Pitch is the black residue of.
 - Coke
 - Coaltar
 - Coal
 - Coal gas
- Peat contains carbon
 - 60%
 - 70%
 - 80%
 - 90%

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Name the gases which are found in coal gas?
- Define catenation. Give its conditions.
- Draw the structural formulae of n-butane and isobutene.
- Define isomerisms. Give examples.
- What is the importance of natural gas?

Q. No.3. How is coal formed? What are the different types of coal? Explain. 5**SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Group-B	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Coal gas is the mixture of.
 - CO and CH₄
 - CO, CH₄, CO₂
 - CO, H₂, CH₂
 - none
- Which of the following is synthetic fiber?
 - Cotton
 - nylon
 - wool
 - silk
- The percentage of carbon in lignite is.
 - 60%
 - 90%
 - 80%
 - None of these
- The Superior quality of coal is
 - Peat
 - lignite
 - Anthracite
 - none
- Coal Tar is the mixture of compounds more than.
 - 200
 - 300
 - 500
 - 100

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is ammoniacal liquor?
- How coal is formed?
- What are aromatic compounds? Give example.
- What is isomerism? Give the structure of isomers of butane.
- Which types of compounds are synthesized by plants?

Q. No.3. Write down the characteristics of organic compounds?**5**

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Test: Chemistry	Group-A	Class: 10th A
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which of the following is hardest coal?
 - Lignite
 - peat
 - Anthracite
 - none
- Main components of natural gas is
 - Methane
 - Propane
 - butane
 - propane
- Which of the following does not contain protein
 - Pulses
 - potatoes
 - eggs
 - beans
- Pitch is the black residue of.
 - Coke
 - Coaltar
 - Coal
 - Coal gas
- Peat contains carbon
 - 60%
 - 70%
 - 80%
 - 90%

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Name the gases which are found in coal gas?
- What are aromatic compounds?
- Draw the structural formulae of n-butane and isobutene.
- Define heterocyclic compounds. Give examples.
- What is the importance of natural gas?

Q. No.3. How is coal formed? What are the different types of coal? Explain.**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Group-B	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Coal gas is the mixture of.
 - CO and CH₄
 - CO, CH₄, CO₂
 - CO, H₂, CH₂
 - none
- Which of the following is synthetic fiber?
 - Cotton
 - nylon
 - wool
 - silk
- The percentage of carbon in lignite is.
 - 60%
 - 90%
 - 80%
 - None of these
- The Superior quality of coal is
 - Peat
 - lignite
 - Anthracite
 - none
- Coal Tar is the mixture of compounds more than.
 - 200
 - 300
 - 500
 - 100

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is ammoniacal liquor?
- How coal is formed?
- What are aromatic compounds? Give example.
- What is isomerism? Give the structure of isomers of butane.
- Which types of compounds are synthesized by plants?

Q. No.3. Write down the characteristics of organic compounds?**5**

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Test: Chemistry	Group-A	Class: 10th A
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Which of the following is hardest coal?
 - Lignite
 - peat
 - Anthracite
 - none
- Main components of natural gas is
 - Methane
 - Propane
 - butane
 - propane
- Which of the following does not contain protein
 - Pulses
 - potatoes
 - eggs
 - beans
- Pitch is the black residue of.
 - Coke
 - Coaltar
 - Coal
 - Coal gas
- Peat contains carbon
 - 60%
 - 70%
 - 80%
 - 90%

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Name the gases which are found in coal gas?
- Define catenation. Give its conditions.
- Draw the structural formulae of n-butane and isobutene.
- Define isomerisms. Give examples.
- What is the importance of natural gas?

Q. No.3. How is coal formed? What are the different types of coal? Explain.**5****SEDIINFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Group-B	Class: 10th
Time: 40 Min.	Date:	Max. Marks: 20
Name: _____		Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Coal gas is the mixture of.
 - CO and CH₄
 - CO, CH₄, CO₂
 - CO, H₂, CH₂
 - none
- Which of the following is synthetic fiber?
 - Cotton
 - nylon
 - wool
 - silk
- The percentage of carbon in lignite is.
 - 60%
 - 90%
 - 80%
 - None of these
- The Superior quality of coal is
 - Peat
 - lignite
 - Anthracite
 - none
- Coal Tar is the mixture of compounds more than.
 - 200
 - 300
 - 500
 - 100

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is ammoniacal liquor?
- How coal is formed?
- What are aromatic compounds? Give example.
- What is isomerism? Give the structure of isomers of butane.
- Which types of compounds are synthesized by plants?

Q. No.3. Write down the characteristics of organic compounds?**5**

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Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: A

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- The functional group $-\text{COOH}$ found in.
 - alcohol
 - carboxylic
 - aldehydes
 - ketones
- Which contains triple bond?
 - pentane
 - ethylene
 - acetylene
 - none
- Which one of following is not a fossil fuel?
 - coal
 - natural gas
 - biogas
 - petroleum
- Which one of the following does not contain proteins?
 - pulses
 - potatoes
 - beans
 - eggs
- Main component of natural gas is
 - methane
 - propane
 - butane
 - propane

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define functional group. Give Example.
- What is the difference between n-propyl and iso propyl.
- What is the difference between aldehyde and ketones.
- What are identification test for ketonic group?
- What is the basic unit of carbohydrates and how it is synthesized?

Q. No.3. Write down the used of organic compounds.**5****SEDiNFO.NET SCIENCE ACADEMY**sedinfo.net@gmail.com

Test: Chemistry	Class: 10th
Time: 40 Min.	Max. Marks: 20
Name: _____	Date: _____
	Section: A

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Coal tar contains compounds
 - benzene
 - phenol
 - toluene
 - all
- Which contains triple bond?
 - pentane
 - ethylene
 - acetylene
 - none
- Who was prepared acetic acid in laboratory?
 - Berzelius
 - wohler
 - kolble
 - Dalton
- Which one of the following does not contain proteins?
 - pulses
 - potatoes
 - beans
 - eggs
- Main component of natural gas is
 - methane
 - propane
 - butane
 - propane

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- How are alkyl radicals formed? Explain with example.
- What is the difference between n-propyl and iso propyl.
- What is an ester group? Write the formula of ethyl acetate.
- What are identification test for ketonic group?
- What is the basic unit of carbohydrates and how it is synthesized?

Q. No.3. Write the characteristics properties of homologous series.**5**

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Paper: CHEMISTRY

Time: 15 Min.

Name: _____

Class: 10th

Max. Marks: 12

Dated: _____

OBJECTIVE TYPE(2nd 50%)

Note: Give answers to the questions on the objective answer sheet provided. Four possible answers A, B, C or D are given for each question. Fill the circle (A, B, C or D) which you consider correct with marker or pen.

Q.1 [1 × 12= 12]

1. Which is not a heavy metal?
(a) Cadmium (b) Lead (c) Zinc (d) Mercury
2. Which of the following is a triglyceride?
(a) Carbohydrates (b) Proteins (c) Lipids (d) Vitamins
3. Percentage of carbon dioxide in dry air is
(a) 21% (b) 0.93% (c) 0.03% (d) None of these
4. Thousands of the amino acid polymerize to form
(a) carbohydrates (b) proteins (c) lipids (d) vitamins
5. The percentage of Sunlight absorbed by atmosphere gases is
(a) 2% (b) 10% (c) 18% (d) 25%
6. Which gas acts as a glass wall of a green house?
(a) Oxygen (b) Carbon Dioxides (c) Sulphur Dioxide (d) Hydrogen
7. Which one of the following salts makes the water permanently hard?
(a) Na₂CO₃ (b) NaHCO₃ (c) Ca(HCO₃)₂ (d) CaSO₄
8. Which one of the following vitamins is water soluble?
(a) vitamin A (b) vitamin E (c) vitamin C (d) vitamin D
9. Water has a maximum density at
(a) 10°C (b) 0°C (c) 4°C (d) 100°C
10. Which contains sufficient amount of metal
(a) Mineral (b) Ores (c) Rocks (d) Soil
11. The process by which atmospheric nitrogen is turned into nitrates in the soil is called:
(a) nitration (b) fixing (c) oxidation (d) redaction
12. Which one of the followings is not a fraction of petroleum?
(a) kerosene oil (b) diesel oil (c) alcohol (d) petrol