

**SEDiNFO.NET Academy**Whatsapp Contact No. [+923156673963](https://www.whatsapp.com/channel/0029156673963)

Paper: Chemistry

Class: 9<sup>th</sup>

Date:

Group-

- =====
- Q1. Encircle the correct option** **05**
- (i) **Heat of vaporation of water is**  
 (a) 40.7 KJ/mol (b) 60.7KJ/mol (c) 100.0 KJ/mol (d) 35.5KJ/mol
- (ii) **How many pascals in one atm. pressure?**  
 (a)101325 (b)10325 (c)106075 (d)10523
- (iii) **Which are of the following example of amorphous solid.**  
 (a) daimond (b)glucose (c) NaCl (d) glass
- (iv) **What is the density of water?**  
 (a) 1.0g cm<sup>-1</sup> (b)1.5g cm<sup>-1</sup> (c) 0.001g cm<sup>-1</sup> (d)0.1g cm<sup>-1</sup>
- (v) **How many tines liquids are denser then gases**  
 (a) 100 times (b)1000 times (c) 10.00 times (d) 100.00 times
- Q2. Give short answers:** **10**
- (i) Define evaporation give example.  
 (ii) Give the difference between amorphous and crystalline solids.  
 (iii) Why density of solids is higher then gases?  
 (iv) Evaporation causes cooling. Explain.  
 (v) Why evaporatin increases with increase in temperature.
- Q3. Write a note on Allotropy with examples.** **05**

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- Q1. Encircle the correct option** **05**
- (i) **At 0<sup>o</sup>c vapour pressure of diethyl ether is**  
 (a) 200 mmHg (b) 25mmHg (c) 5mmHg (d) 150mmHg
- (ii) **Which are of the following is not amorphous?**  
 (a)rubber (b) plastic (c) glass (d) Glucose
- (iii) **Density of gas increase, when its:**  
 (a) Temp increased (b) pressure is increased (c) volume is kept constant (d) none of these
- (iv) **Solid are rigid and \_\_\_\_\_ then liquids**  
 (a) denser (b)lighter (c)faster (d)Both A and B
- (v) **Densities of gases are expressed in terms of:**  
 (a) mg cm<sup>-3</sup> (b)g cm<sup>-3</sup> (c) g dm<sup>-3</sup> (d) kg dm<sup>-3</sup>
- Q2. Give short answers:** **10**
- (i) Define Boiling point. Give examples.  
 (ii) Define transition temperature with one example.  
 (iii) Why are liquids mobile?  
 (iv) Why rate of vapour pressure increases with increase in temperature?  
 (v) Define allotropy. Also write the examples of allotropy.
- Q3. Describe the phenomenon of evaporation along with factors.** **05**

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Date: 19-9-2014Class: 9<sup>th</sup> A  
Group-

- =====
- Q1. Encircle the correct option** **5×1=5**
- (i) **Which one of these have lowest boiling point.**  
(a) Ether (b) alcohol (c) Water (d) acids
- (ii) **The boiling point does not depends upon.**  
(a) Nature of liquid (b) intermolecular force (c) External pressure (d) density
- (iii) **Density of H<sub>2</sub>O is**  
(a) 0.001 gcm<sup>-3</sup> (b) 1 gcm<sup>-3</sup> (c) 0.0001 gcm<sup>-3</sup> (d) 0.01 gcm<sup>-3</sup>
- (iv) **The diamond having regularly arranged molecules are**  
(a) Amorphous (b) metallic (c) covalent (d) crystalline
- (v) **White Sn is in shape**  
(a) Cubic (b) pyramidal (c) Trigonal (d) Tetragonal
- Q2. Give short answers:** **5×2=10**
- (i) Define transition temperature.
- (ii) Why the boiling point of diethyl ether increases rapidly?
- (iii) Gases have low density than solids. Justify?
- (iv) Define vapour pressure?
- (v) How nature of liquid affects boiling point?
- Q3.a) Write a note on types of solids.** **03**
- b) Discuss the factors affecting diffusion.** **02**

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Group-

- =====
- Q1. Encircle the correct option** **5×1=5**
- (i) **Density of air is**  
(a) 0.001 gcm<sup>-3</sup> (b) 0.0001 gcm<sup>-3</sup> (c) 0.01 gcm<sup>-3</sup> (d) 0.1 gcm<sup>-3</sup>
- (ii) **Diffusion is not effected by**  
(a) size of molecules (b) Temperature (c) Density (d) Intermolecular
- (iii) **The rate of diffusion of gases is**  
(a) Slow (b) fast (c) both a and b (d) none of these
- (iv) **The shape of S<sub>8</sub> is**  
(a) rhombic (b) monoclinic (c) both a and b (d) cubic
- (v) **Density of Iron is**  
(a) 2.70 gcm<sup>-3</sup> (b) 19.3 gcm<sup>-3</sup> (c) 7.86 gcm<sup>-3</sup> (d) 1.0 gcm<sup>-3</sup>
- Q2. Give short answers:** **5×2=10**
- (i) Define melting point?
- (ii) Water has high B.P than other, why?
- (iii) How shape of molecules affecting on diffusion?
- (iv) Define density.
- (v) What are crystalline solids?
- Q3.a) Discuss the factors affecting the boiling point?** **05**