

Test: Chemistry	Class: 9th
Time: 40 Min.	Max. Marks: 20
Name: _____	Section: _____
Date: _____	

Q. No.1. Choose the Correct Answer.**1 x 5 = 5**

- Densities of gases are expressed in term of
 - mg cm^{-3}
 - g cm^{-3}
 - mg dm^{-3}
 - kg dm^{-3}
- One atmospheric pressure is equal to
 - 101325 pa
 - 10325 Pa
 - 10675 Pa
 - 105123 Pa
- Which one of the following gas diffuses faster?
 - Hydrogen
 - Helium
 - fluorine
 - chlorine
- Density of oxygen at 20°C is
 - 1.4 g dm^{-3}
 - 1.5 g dm^{-3}
 - 1.4 g dm^{-3}
 - 1.7 g dm^{-3}
- Which of the following state has strong intermolecular forces?
 - solids
 - liquid
 - gas
 - none of these

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define diffusion and effusion.
- Why the densities of gases are lower than solid and liquids?
- Why the gases are compressible?
- Convert 1.25 atm to Pascals?
- Convert 172 K to °C

Q. No.3. Define Boyle's law Explain with an example.**5**

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(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why the gases are compressible?
- Convert 172 K to °C.
- Define diffusion and effusion.
- What is absolute zero?
- Why the densities of gases are lower than solid and liquids?

Q. No.3. Define Charle's law Explain with an example.**5**