

<b>Test: Chemistry</b> <b>Time: 40 Min.</b> <b>Name:</b> _____	<b>Date: 09-07-2014</b>	<b>Class: 9<sup>th</sup></b> <b>Max. Marks: 20</b> <b>Section:</b>
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**Q. No.1. Choose the Correct Answer.**

1 x 5 = 5

- Which molecule contain a double covalent bond?  
a) CH<sub>4</sub>      b) C<sub>2</sub>H<sub>4</sub>      c) C<sub>2</sub>H<sub>4</sub>      d) none
- How many valence electrons in sodium ion?  
a) 01      b) 02      c) 03      d) 08
- How many lone pairs are present on nitrogen in ammonia molecule?  
a) one      b) two      c) three      d) four
- Hydrogen and helium follows.  
a) octet rule      b) duplet rule      c) triplet rule      d) none
- The formation of ionic bond between two ion is due to  
a) Hydrogen bonding      b) metallic bond  
c) electrostatic forces      d) none

(SUBJECTIVE TYPE)

**Q. No.2. Give Short Answers of the following questions.**

2 x 5 = 10

- Define double covalent bond. Give example.
- Why does sodium form chemical bond with chlorine?
- Write the electronic configuration of chloride ion.
- Why noble gases are not reactive?
- Define ionic bond. Give example.

**Q. No.3. Define coordinate covalent bond. Explain with examples.**

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**Q. No.1. Choose the Correct Answer.**

1 x 5 = 5

- Which molecule contain a double covalent bond?  
a) CH<sub>4</sub>      b) C<sub>2</sub>H<sub>4</sub>      c) C<sub>2</sub>H<sub>4</sub>      d) none
- How many valence electrons in sodium ion?  
a) 01      b) 02      c) 03      d) 08
- How many lone pairs are present on nitrogen in ammonia molecule?  
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- Hydrogen and helium follows.  
a) octet rule      b) duplet rule      c) triplet rule      d) none
- The formation of ionic bond between two ion is due to  
a) Hydrogen bonding      b) metallic bond  
c) electrostatic forces      d) none

(SUBJECTIVE TYPE)

**Q. No.2. Give Short Answers of the following questions.**

2 x 5 = 10

- Define double covalent bond. Give example.
- Why does sodium form chemical bond with chlorine?
- Write the electronic configuration of chloride ion.
- Why noble gases are not reactive?
- Define Polar and non polar bond.

**Q. No.3. Define ionic bond. Explain with examples.**

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