

Test: Chemistry Time: 40 Min. Name: _____	Date: 09-07-2014	Class: 9th Max. Marks: 20 Section:
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Q. No.1. Choose the Correct Answer.

1 x 5 = 5

- Which molecule contain a double covalent bond?
a) CH₄ b) C₂H₄ c) C₂H₄ d) none
- How many valence electrons in sodium ion?
a) 01 b) 02 c) 03 d) 08
- How many lone pairs are present on nitrogen in ammonia molecule?
a) one b) two c) three d) four
- Hydrogen and helium follows.
a) octet rule b) duplet rule c) triplet rule d) none
- The formation of ionic bond between two ion is due to
a) Hydrogen bonding b) metallic bond
c) electrostatic forces d) none

(SUBJECTIVE TYPE)

Q. No.2. Give Short Answers of the following questions.

2 x 5 = 10

- Define double covalent bond. Give example.
- Why does sodium form chemical bond with chlorine?
- Write the electronic configuration of chloride ion.
- Why noble gases are not reactive?
- Define ionic bond. Give example.

Q. No.3. Define coordinate covalent bond. Explain with examples.

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(SUBJECTIVE TYPE)

Q. No.2. Give Short Answers of the following questions.

2 x 5 = 10

- Define double covalent bond. Give example.
- Why does sodium form chemical bond with chlorine?
- Write the electronic configuration of chloride ion.
- Why noble gases are not reactive?
- Define Polar and non polar bond.

Q. No.3. Define ionic bond. Explain with examples.

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