

**SEDiNFO.NET Academy**Whatsapp Contact No. [+923156673963](https://www.whatsapp.com/business/profile/923156673963)

<b>Test: Chemistry</b>	<b>Class: 9<sup>st</sup></b>
<b>Time: 40 Min.</b>	<b>Max. Marks: 20</b>
<b>Name: _____</b>	<b>Section: _____</b>
<b>Group- _____</b>	
<b>Date: _____</b>	

**Q1. Answer the following Short Question.****05****1. The first ionization energy of sodium atom is.**

- (a) 495.8 KJ mol<sup>-1</sup> (b) -495.8 KJ mol<sup>-1</sup>  
 (c) 328 KJ mol<sup>-1</sup> (d) None of these

**2. Which element has high electronegativity?**

- (a) Chlorine (b) bromine (c) iodine (d) flourine

**3. The unit of atomic size is.**

- (a) nm (b) pm (c) KJmol<sup>-1</sup> (d) a & b Both

**4. Along the periods, which one of following decreases.**

- (a) atomic radius (b) electronegativity (c) ionization energy (d) electron effimity

**5. Transition elements are.**

- (a) all gases (b) all metals (c) all non metals (d) all metallds

**Q. No.2. Give Answers of the following short questions.****10**

1. What is trend of ionization energy in periods?
2. Define shielding effect.
3. Define elctronegativity. Give example.
4. Why electron affinity increase in periods?
5. Why shielding effect makes cation formation easy?

**Q. No.3.Long Question.**

- 1) Define atomic size. Write its trends in periodic table.

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