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Test: Chemistry	Ch # 1	Class: 9th
Time: 40 Min.	Group	Max. Marks: 20
Name: _____	Date:	Section:

Q. No.1. Choose the Correct Answer.**1 x 5 = 05**

- Percentage of argon in nature is
a) 87% b) 47% c) 0.9% d) 21%
- Example of poly atomic molecule.
a) CH₄ b) H₂SO₄ c) C₆H₁₂O₆ d) All of these
- Empirical formula of acetic acid (CH₃COOH) is
a) CHO b) CH₂O c) CHO₂ d) none of these
- How many numbers of moles in 8g of CO₂?
a) 0.18 b) 0.15 c) 0.24 d) 0.21
- The molar mass of H₂SO₄ is
a) 98g b) 98amu c) 9.8g d) 9.8 amu

(SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5=10**

- What is the relative atomic mass? How it is related to gram? 2. Define analytical chemistry. Give its scope.
- How does homogenous mixture differ from heterogeneous mixture?
 - Define empirical formula. Give example. 5. Define mole and Avogadro's number.

Q. No.3. a) Differentiate between mixture and compound.**5****StudyNowPK.COM Academy**Whatsapp Contact No. [+923156673963](https://www.whatsapp.com/channel/0029156673963)

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SUBJECTIVE TYPE)**Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- What is the relative atomic mass? 2. Define analytical chemistry. Give its scope.
- How does homogenous mixture differ from heterogeneous mixture?
- Define empirical formula. Give example. 5. Find the molecular mass of H₂SO₄.

Q. No.3. a) Differentiate between mixture and compound.**3****b) Differentiate between atom and ion.****2**