

<b>Test: Chemistry</b>	<b>Class: 9<sup>th</sup></b>
<b>Time: 40 Min.</b>	<b>Max. Marks: 20</b>
<b>Name: _____</b>	<b>Section: _____</b>
<b>Date: _____</b>	

**Q. No.1. Choose the Correct Answer.****1 x 5 = 5**

- Densities of gases are expressed in term of
  - $\text{mg cm}^{-3}$
  - $\text{g cm}^{-3}$
  - $\text{mg dm}^{-3}$
  - $\text{kg dm}^{-3}$
- One atmospheric pressure is equal to
  - 101325 pa
  - 10325 Pa
  - 10675 Pa
  - 105123 Pa
- Which one of the following gas diffuses faster?
  - Hydrogen
  - Helium
  - fluorine
  - chlorine
- Density of oxygen at 20°C is
  - $1.4 \text{ g dm}^{-3}$
  - $1.5 \text{ g dm}^{-3}$
  - $1.4 \text{ g dm}^{-3}$
  - $1.7 \text{ g dm}^{-3}$
- Which of the following state has strong intermolecular forces?
  - solids
  - liquid
  - gas
  - none of these

**(SUBJECTIVE TYPE)****Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Define diffusion and effusion.
- Why the densities of gases are lower than solid and liquids?
- Why the gases are compressible?
- Convert 1.25 atm to Pascals?
- Convert 172 K to °C

**Q. No.3. Define Boyle's law Explain with an example.****5**

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  - solids
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**(SUBJECTIVE TYPE)****Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why the gases are compressible?
- Convert 172 K to °C.
- Define diffusion and effusion.
- What is absolute zero?
- Why the densities of gases are lower than solid and liquids?

**Q. No.3. Define Charle's law Explain with an example.****5**