

Test: Chemistry		Class: 9 <sup>th</sup>
Time: 40 Min.		Max. Marks: 20
Name: _____	Date: _____	Section: _____

**Q. No.1. Choose the Correct Answer.****1 x 5 = 5**

- Hydrogen and helium follows
  - Octet rule
  - Duplet rule
  - Triplet rule
  - none of them
- Metals have the tendency to lose electrons due to
  - High ionization energy
  - Low electron affinity
  - Low ionization energy
  - None of them
- Intermolecular forces are collectively known as
  - Vander Waals forces
  - Electrostatic forces
  - Adhesive forces
  - Dipole-dipole forces
- Malleability is the property by virtue of which a metal can be drawn into
  - Sheets
  - Wires
  - Rods
  - Plates
- Epoxy adhesives are stable to heat up to a temperature of
  - 177 °C
  - 225 °C
  - 320 °C
  - 135 °C

**(SUBJECTIVE TYPE)****Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why ionic compounds have high melting and boiling points?
- How metallic bond is formed?
- How dipole-dipole interactions are found in HCl molecule?
- Why a dipole develops in a molecule?
- Why the covalent compounds of bigger size molecules have high melting points?

**Q. No.3. Write the characteristics properties of ionic compounds.****05**

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**(SUBJECTIVE TYPE)****Q. No.2. Give Short Answers of the following questions.****2 x 5 = 10**

- Why the covalent compounds of bigger size molecules have high melting points?
- Why a dipole develops in a molecule?
- How dipole-dipole interactions are found in HCl molecule?
- Why ionic compounds have high melting and boiling points?
- How metallic bond is formed?

**Q. No.3. Write the characteristics properties of ionic compounds.****05**