

**SEDiNFO.NET Academy**

Whatsapp Contact No. [+923156673963](https://www.whatsapp.com/business/profile/923156673963)

Marks: \_\_\_\_\_  
Name: \_\_\_\_\_

Group A

Time: \_\_\_\_\_

Class: 9<sup>th</sup>

Date: \_\_\_\_\_

Subject: Chemistry

**Q.1 Choose the Correct Option**

- 1) Ionic solids and polar covalent compounds are soluble in.  
a) Benzene                      b) ether                      c) Water                      d) Petrol
- 2) The compounds soluble in water are.  
a) KCl                      b) Na<sub>2</sub>CO<sub>3</sub>                      c) CuSO<sub>4</sub>                      d) All of these
- 3) Which of the following shows tyndall effect.  
a) Albumin                      b) milk                      c) points                      d) both a and b
- 4) Which one is suspension?  
a) Blood                      b) tooth past                      c) Ink                      d)chalk in water

**Q.2 Short Answers.**

1. Define solubility?
2. Write any two differences between solution and suspension?
3. 10cm<sup>3</sup> of 0.01M KMnO<sub>4</sub> has been dilated to 100cm<sup>3</sup>. Find out the molarity of this solution.
4. How does the solubility of a salt increase with the increase in temperature .Give example
5. What is the effect of non-polar substances on polar solvents?

**Q.3 Long Question.**

- a. Give the comparison between colloids and suspensions.

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Subject: Chemistry

**Q.1 Choose the Correct Option**

- 1) Which one is suspension?  
a) Blood                      b) Tooth paste                      c) Ink                      d) Chalk in water
- 2) Which one of the following is heterogeneous mixture?  
a) Milk                      b) Ink  
c) Milk of magnesia                      d) Sugar solution
- 3) The solubility of which salt decrease with the increase in temperature.  
a) KNO<sub>3</sub>                      b) NaNO<sub>3</sub>                      c) Li<sub>2</sub>SO<sub>4</sub>                      d) KCl
- 4) Tyndall effect is show by.  
a) Sugar solution                      b) points                      c) jelly                      d) Chalk solution

**Q.2 Short Answers.**

1. Name the factors which affect the solubility of solute?
2. What do you mean like 'dissolves like' explain with example.
3. Write any two differences between colloids and suspension.
4. How much NaOH is required to prepare its 500cm<sup>3</sup> of 0.4M solution?
5. Why suspensions and solutions do not show Tyndall effect while colloids do?

**Q.3 Long Question.**

- a. Explain the effect of temperature on solubility.