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Test: Chemistry		**Class: 9th**
Time: 40 Min.	**Group**	**Max. Marks: 20**
Name: _____	**Date:**	**Section:**

Q. No.1. Choose the Correct Answer. **1 x 5 = 5**

1. Which one of the following halogen has lowest electro negativity?
a) fluorine b) chlorine c) bromine d) iodine

2. Transition elements are.
a) all gases b) all metals c) all liquids d) none

3. The unit of electro negativity is.
a) Kj mol^{-1} b) Nm^{-2} c) nm d) no unit

4. 1st ionization energy of sodium atom is.
a) +496 $kjmol^{-1}$ b) -328 $kjmol^{-1}$ c) 180 $kjmol^{-1}$ d) none

5. Electro negativity of fluorine is.
a) 4.0 b) 3.5 c) 3.2 d) 2.2

(SUBJECTIVE TYPE)

Q. No.2. Give Short Answers of the following questions. **2 x 5 = 10**

1. Define shielding effect.
2. Why atomic size increases in group and decreases in period?
3. Define electro negativity. Give example.
4. Why shielding effect make cation formation easy?
5. Define electron affinity. Give its units.

Q. No.3. Define ionization energy. Give its trends. **5**

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